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Modern Chemistry • CHAPTER 7

## HOMework 7-1

(pp. 261–266)

### VOCABULARY

Write *true* or *false* for each statement.

1. A **chemical formula** indicates the relative number of molecules of each kind in a chemical compound.  
\_\_\_\_\_
2. **Monoatomic ions** are ions formed from a single atom. \_\_\_\_\_
3. **Monoatomic anions** of an element are identified by the element's name. \_\_\_\_\_
4. A **binary compound** is made of two or more different elements. \_\_\_\_\_
5. The **Stock system** is a system for naming chemical elements and their ions. \_\_\_\_\_

### SKILL BUILDER

Explain what each chemical formula tells you about the composition of the compound. The first one is done for you.



The subscript 3 refers to the number of oxygen atoms. There are one atom of sodium and one atom of nitrogen for every three atoms of oxygen.



### STANDARDIZED TEST PREP

Circle the letter of the best answer.

1. Which statement is true?
  - a. All main-group elements are able to form ions.
  - b. The total numbers of positive charges and negative charges must be equal in a binary compound.
  - c. The nonmetals of groups 15, 16, and 17 lose electrons to form cations.
  - d. Elements in the *d* block form ions of only one charge.
2. Why do many main-group elements lose or gain electrons to form ions?
  - a. The elements are unstable and therefore lose electrons.
  - b. Their tendency is to maximize potential energy.
  - c. The elements form covalent bonds, sharing their electrons with other atoms.
  - d. Their tendency is to form a complete outermost octet in a noble-gas configuration.

# Holt Modern Chemistry Chapter 7 Test Answers

**L Reisser**



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