

The Activity Series

Use with Chapter 11
Section 11.2

1. The ends of the following pairs of elements, indicating the one that could replace the other in the given compound:

- | | |
|------------------------|--------------------|
| a. calcium, tin | e. zinc, silver |
| b. sodium, barium | f. sodium, calcium |
| c. aluminum, potassium | g. silver, lead |
| d. iron, copper | |

2. The ends of the following reactions are the active series for aluminum, identify the product (metal compound) and, if appropriate, give the balanced equation. If the reaction does not occur, indicate the formula for the product of the reaction. (Note: In an active metal displaces the hydrogen in acids; the formula of the acid is not given.)

- Lead + Sulfuric acid → $PbSO_4 + H_2$
- Lead + Nitric acid → no rxn
- Copper + Nitric acid → $Cu(NO_3)_2 + H_2$
- Copper + Sulfuric acid → no rxn
- Silver + Nitric acid → $AgNO_3 + H_2$
- Silver + Sulfuric acid → no rxn

3. Magnesium metal reacts with various acids. Does silver metal (Ag) react in the following reactions? Write the balanced equation if the reaction occurs, otherwise, explain the non-reaction.

Ag + hydrochloric acid → no rxn Ag + sulfuric acid → no rxn Ag + nitric acid → no rxn
 Ag + acetic acid → no rxn Ag + hydrofluoric acid → no rxn

4. Write the balanced equation for the reaction of silver metal with the following acids. If no reaction occurs, explain why.

Ag + hydrochloric acid → no rxn Ag + nitric acid → no rxn Ag + acetic acid → no rxn

5. The ends of the reactions below are the active series for aluminum, identify the product (metal compound) and, if appropriate, write the balanced equation. If the reaction does not occur, indicate the formula for the product of the reaction. (Note: In an active metal displaces the hydrogen in acids; the formula of the acid is not given.)

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