

Stoichiometry Worksheet # 1

1) Glucose is used as a source of energy by the human body. The overall reaction in the body is



Calculate the number of moles of oxygen needed to oxidize 12.5 g of $\text{C}_6\text{H}_{12}\text{O}_6$.

2) Ammonia is synthesized from hydrogen and nitrogen according to the following equation.



If an excess of nitrogen is reacted with 3.41 moles of hydrogen gas, how many grams of ammonia (NH_3) can be produced?

3) Assume that in the decomposition of potassium chlorate, KClO_3 ,



80.5 moles of O_2 form. How many grams of potassium chlorite, the other product would be formed?

4) In a single displacement reaction,



9.23 moles of aluminum react with excess hydrochloric acid. How many moles of hydrogen will be produced?

5) The compound "Cisplatin" $\text{PtCl}_2(\text{NH}_3)_2$, has been found to be effective in treating some types of cancer. It can be synthesized using the following reaction



A. How many moles of "cisplatin" can be produced from 2.50 g K_2PtCl_4 ?

B. How many moles of NH_3 would be needed if 3.8 moles of K_2PtCl_4 were present?

Pearson Education Chapter 12 Stoichiometry Answer Key

Jessica J Manson



Pearson Education Chapter 12 Stoichiometry Answer Key:

The book delves into Pearson Education Chapter 12 Stoichiometry Answer Key. Pearson Education Chapter 12 Stoichiometry Answer Key is a crucial topic that must be grasped by everyone, from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Pearson Education Chapter 12 Stoichiometry Answer Key, encompassing both the fundamentals and more intricate discussions.

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