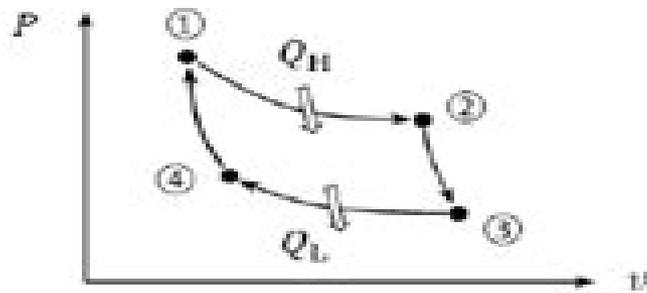# Written Problem

The Carnot cycle is an idealization that involves four processes as shown (see also Figure 6–37 in textbook). As each of the processes is reversible, the entire cycle is also reversible.

$① → ②$: Isothermal expansion at $T = T_H$

$② → ③$: Adiabatic expansion (no heat transfer)

$③ → ④$: Isothermal compression at $T = T_L$

$④ → ①$: Adiabatic compression (no heat transfer)

(a) Suppose the Carnot cycle is carried out in a closed system with an ideal gas as working fluid. Show that

$$Q_H = mRT_H \ln(V_2/V_1)$$

where m is the mass of the ideal gas. (Recall that the internal energy of an ideal gas depends *only* on the temperature T.)

You are now provided with more specifics: The ideal gas is 2 g of air. The Carnot cycle operates between $T_H = 500$ K and $T_L = 300$ K, with $V_2 = 2V_1$.

(b) What is the amount of heat Q_L (kJ) that is transferred to the cold reservoir in every cycle?

(c) Determine W_{net} (kJ), the net amount of work that is produced per cycle.

Note that W_{net} can be obtained by summing the $\int P\,dV$ contribution from each of the four processes. This approach, however, is not recommended as there are much easier ways to get to the answer.

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