

Solutions

- Solubility of gas in liquid : (Henry's law)

② $K_H \propto \frac{1}{\text{Solubility}}$

P.P $\propto$ solubility

P.P $\propto$ mole fraction of gas

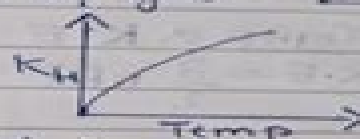
$$PP = K_H \times x_{\text{gas}}$$

K_H = Henry's constant of gas

Depends on nature of gas

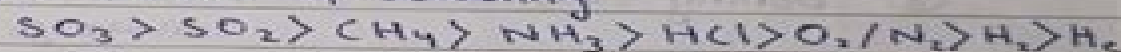


① $K_H \propto \text{Temp}$



- Factors affecting solubility of Gas:

- ① Nature of gas : order of solubility



- ② Nature of solvent $\propto$ Dielectric constant $\propto$ Polarity

- ③ $\text{Temp} \propto \frac{1}{\text{Solubility}}$

④ $P.P = K_H \times \text{solubility}$
 $PP \propto \text{solubility}$

• Limitations of Henry's law

- ① $P = \text{low}$ $T = \text{low}$

- ② Gas dissolved in liq

- ③ Ideal gas Behaviour

- ④ V.V small amount of gas soluble in liq

- Solubility of liquid in liquid :

- ① Evaporation : liquid $\rightarrow$ Vapour

- ② condensation : vapour $\rightarrow$ liquid

- ③ At equilibrium pressure exerted by vapour on the surface area of liquid when both liquid and vapour both are at equilibrium.

$$\left. \begin{array}{l} \pi_{\text{eva}} = \pi_{\text{cond}} \\ \text{(at equilibrium state)} \end{array} \right\}$$

↓
Vapour pressure $\rightarrow$ Depends only on the Temperature

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Table of Contents Iit Jee Notes Pdf Webxmedia

1. Understanding the eBook Iit Jee Notes Pdf Webxmedia
 - The Rise of Digital Reading Iit Jee Notes Pdf Webxmedia
 - Advantages of eBooks Over Traditional Books
2. Identifying Iit Jee Notes Pdf Webxmedia
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Iit Jee Notes Pdf Webxmedia
 - User-Friendly Interface
4. Exploring eBook Recommendations from Iit Jee Notes Pdf Webxmedia
 - Personalized Recommendations
 - Iit Jee Notes Pdf Webxmedia User Reviews and Ratings

- Iit Jee Notes Pdf Webxmedia and Bestseller Lists
- 5. Accessing Iit Jee Notes Pdf Webxmedia Free and Paid eBooks
 - Iit Jee Notes Pdf Webxmedia Public Domain eBooks
 - Iit Jee Notes Pdf Webxmedia eBook Subscription Services
 - Iit Jee Notes Pdf Webxmedia Budget-Friendly Options
- 6. Navigating Iit Jee Notes Pdf Webxmedia eBook Formats
 - ePub, PDF, MOBI, and More
 - Iit Jee Notes Pdf Webxmedia Compatibility with Devices
 - Iit Jee Notes Pdf Webxmedia Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Iit Jee Notes Pdf Webxmedia
 - Highlighting and Note-Taking Iit Jee Notes Pdf Webxmedia
 - Interactive Elements Iit Jee Notes Pdf Webxmedia
- 8. Staying Engaged with Iit Jee Notes Pdf Webxmedia
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Iit Jee Notes Pdf Webxmedia
- 9. Balancing eBooks and Physical Books Iit Jee Notes Pdf Webxmedia
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Iit Jee Notes Pdf Webxmedia
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Iit Jee Notes Pdf Webxmedia
 - Setting Reading Goals Iit Jee Notes Pdf Webxmedia
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Iit Jee Notes Pdf Webxmedia
 - Fact-Checking eBook Content of Iit Jee Notes Pdf Webxmedia
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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