

# WORKSHEET

## Acid-base Titration I

Maximum score: 100 points

Your Name:

Partner:

Lab Section (Day & Time):

1. Attach your annotated plots and label the equivalence point and the half-equivalence points appropriately for the titration curves for HCl, KHP, and one trace for the unknown acid, and record the data from the plots. (30 points)

Acid Trial	pH at start of titration	pK <sub>a</sub> or pH at half-equivalence point	pH at equivalence point	pH at end of plot
HCl	2.2	XXXXXXXXXXXX	5.3	11.75
KHP	4.6	5.1	10.5	11.70
Unknown	2.3	2.9	9.0	11.65

2. Complete the table using data from your lab notebook. Calculate the acid concentration using  $M_a V_a = M_b V_b$ . The unknown acid will require a correction for the dilution. (20 pts)

Acid Trial	Volume acid (mL)	NaOH Concentration (M)	Volume of base at Equivalence point (mL NaOH)	Known Acid Concentration (M)	Calculated Acid Concentration (M)
HCl	55	.1111			
KHP	55	.1111	4.43	.1133	.0089
Diluted Unknowns Trial 1	55	.1111	2.54	.0558	.0051
Diluted Unknowns Trial 2		.1111	4.005	XXXXXXXXXX	.0081
				XXXXXXXXXX	

Will not do second trial for unknown

# Acid Base Titration Chemistry If8766 Answer Key

**NARAYAN CHANGDER**



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