

# 17 • Acid-Base Equilibria

## PRACTICE TEST

- What is the  $[H^+]$  when  $[OH^-] = 8.1 \times 10^{-9}$ ?  
 a)  $8.1 \times 10^{-5} M$       d)  $3.6 \times 10^{-8} M$   
 b)  $1.0 \times 10^{-7} M$       e)  $8.1 \times 10^{-2} M$   
 c)  $1.2 \times 10^{-6} M$   $[H^+] = \frac{K_w}{[OH^-]} = \frac{1 \times 10^{-14}}{8.1 \times 10^{-9}} = 1.2 \times 10^{-6}$
- What is the  $[H^+]$  when  $[OH^-] = 3.3 \times 10^{-9}$ ?  
 a)  $3.0 \times 10^{-6} M$       d)  $6.6 \times 10^{-5} M$   
 b)  $1.0 \times 10^{-7} M$       e)  $3.3 \times 10^{-9} M$   
 c)  $3.3 \times 10^{-5} M$   $[H^+] = \frac{K_w}{[OH^-]} = \frac{1 \times 10^{-14}}{3.3 \times 10^{-9}} = 3.3 \times 10^{-6}$
- What is the  $[H^+]$  in a 0.0025 M HCl solution?  
 a)  $1.0 \times 10^{-3} M$       d)  $3.6 \times 10^{-5} M$   
 b)  $4.0 \times 10^{-12} M$       e) need more info  
 c)  $2.5 \times 10^{-3} M$   $HCl \rightarrow H^+ + Cl^-$   
 $[H^+] = .0025 = 2.5 \times 10^{-3}$
- What is the  $[OH^-]$  in a 0.0050 M HCl solution?  
 a)  $5.0 \times 10^{-3} M$       d)  $6.6 \times 10^{-5} M$   
 b) 1.0 M      e)  $2.0 \times 10^{-12} M$   
 c)  $1.0 \times 10^{-7} M$   $[H^+] = .0050$   
 $[OH^-] = \frac{K_w}{[H^+]} = \frac{1 \times 10^{-14}}{.0050} = 2.0 \times 10^{-12}$
- A solution in which  $[H^+] = 10^{-8}$  has a pH of 8 and is basic.  
 a) 8, acidic      d) -8, neutral  
 b) 6, basic      e) 8, basic  
 c) -6, basic
- What is the pH of a 0.00030 M  $HNO_3$  solution?  
 a) 8.11      d) 4.48  
 b) 3.00      e) none of these  
 c) 3.52  $[H^+] = 3 \times 10^{-4}$   
 $pH = -\log [H^+] = 3.52$
- What is the pH of a 0.0060 M KOH solution?  
 a) 5.12      d) 8.88  
 b) 2.22      e) 7.00  
 c) 11.78  $[OH^-] = 6.0 \times 10^{-3}$   
 $pOH = 2.22$   
 $pH = 14 - pOH = 11.78$

- A sample of lemon juice is found to have a pH of 2.55. What is the  $H^+$  concentration of the juice?  
 $[H^+] = 10^{-pH} = 10^{-2.55}$   
 a) 0.0035 M      d) 0.0080 M  
 b) 0.0028 M      e) 355 M  
 c) 11.6 M
- A sample of milk is found to have a pH of 6.60. What is the  $OH^-$  concentration of the milk?  
 a)  $2.5 \times 10^{-22} M$       d)  $4.0 \times 10^{-8} M$   
 b)  $1.0 \times 10^{-7} M$       e)  $2.5 \times 10^{-7} M$   
 c)  $5.0 \times 10^{-7} M$   $[H^+] = 10^{-6.60} = 2.5 \times 10^{-7}$   
 $[OH^-] = \frac{K_w}{[H^+]}$
- What is the concentration of  $OCl^-$  in a 0.60 M solution of HOCl?  $K_a = 3.1 \times 10^{-8}$ .  
 a)  $1.8 \times 10^{-4} M$       d)  $1.4 \times 10^{-4} M$   
 b)  $7.1 \times 10^{-11} M$       e)  $1.1 \times 10^{-4} M$   
 c) 0.40 M  $[OCl^-] = [H^+] \cdot \frac{K_a \cdot .6}{x^2}$   
~~see scratch paper~~
- What is the pH of a 0.020 M solution of hydrosulfuric acid, a diprotic acid?  
 $K_{a1} = 1.1 \times 10^{-7}$        $K_{a2} = 1.0 \times 10^{-14}$  *only 1st dissociation is signif.*  
 a) 7.00      d) 4.33  $\therefore$  use  $K_{a1}$  only  
 b) 9.67      e) 3.05  
 c) 7.84  
~~see scratch paper~~
- What is the concentration of  $CO_3^{2-}$  in a 0.010 M solution of carbonic acid? The relevant equilibria are:  
 $H_2CO_3 \rightleftharpoons H^+ + HCO_3^-$        $K_{a1} = 4.3 \times 10^{-7}$   
 $HCO_3^- \rightleftharpoons H^+ + CO_3^{2-}$        $K_{a2} = 5.6 \times 10^{-11}$   
 a)  $6.6 \times 10^{-5} M$       d)  $7.5 \times 10^{-7} M$   
 b)  $5.6 \times 10^{-11} M$       e)  $7.9 \times 10^{-7} M$   
 c)  $6.7 \times 10^{-11} M$   
~~From the 1st dissociation,  $[H^+] = [HCO_3^-]$   
 the 2nd dissociation doesn't change this much.  
 $K_{a2} = \frac{[H^+][CO_3^{2-}]}{[HCO_3^-]}$  but  $[H^+] = [HCO_3^-]$  cancel!  
 so...  $K_{a2} = [CO_3^{2-}] = 5.6 \times 10^{-11}$~~

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