

## practice test ch 7 & 8

### Multiple Choice

Identify the choice that best completes the statement or answers the question.

- \_\_\_\_\_ 1. Which of the following best describes the rates of chemical reaction?
- most chemical reactions occur very slowly
  - most chemical reactions occur at moderate rates
  - most chemical reactions occur very rapidly
  - chemical reactions have a wide range of rates, from extremely fast to extremely slow
- \_\_\_\_\_ 2. Which of the following is true of the rates of most chemical reactions?
- the initial rate is cannot be measured
  - the initial rate is faster than the rate later in time
  - the initial rate is the same as the rate later in time
  - the initial rate is slower than the rate later in time
- \_\_\_\_\_ 3. Which of the following is true of effective collisions?
- the number of effective collisions decreases as the temperature is increased
  - the number of effective collisions determines the reaction rate
  - both a. and b.
  - Neither a. nor b.
- \_\_\_\_\_ 4. Which of the following is true of effective collisions?
- the number of effective collisions determines the reaction rate
  - the number of effective collisions increases as the temperature is increased
  - both a. and b.
  - Neither a. nor b.
- \_\_\_\_\_ 5. Which of the following is true?
- decreasing the activation increases the reaction rate
  - the activation energy for an exothermic reaction is negative
  - both a. and b.
  - Neither a. nor b.
- \_\_\_\_\_ 6. In an energy diagram for a chemical reaction what species appears at the highest energy?
- the catalyst
  - the products
  - the reactants
  - the transition state
- \_\_\_\_\_ 7. Given that the reaction  $2 \text{H}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2 \text{H}_2\text{O}(\text{g})$  is exothermic, which of the following is true of the reaction  $2 \text{H}_2\text{O}(\text{g}) \rightarrow 2 \text{H}_2(\text{g}) + \text{O}_2(\text{g})$ ?
- its activation energy is lower than that of  $2 \text{H}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2 \text{H}_2\text{O}(\text{g})$
  - its activation energy is the same as that of  $2 \text{H}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2 \text{H}_2\text{O}(\text{g})$
  - its activation energy is higher than that of c. its activation energy is the same as that of  $2 \text{H}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2 \text{H}_2\text{O}(\text{g})$
  - there is no relationship between its activation energy and that of  $2 \text{H}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2 \text{H}_2\text{O}(\text{g})$
- \_\_\_\_\_ 8. In an energy diagram for an endothermic chemical reaction which of the following is true?
- the energy of the products is lower than that of the reactants
  - the energy of the transition state is higher than that of the reactants
  - both a. and b.
  - Neither a. nor b.
- \_\_\_\_\_ 9. Consider a reaction such as  $\text{A}_2(\text{g}) + \text{B}_2(\text{g}) \rightarrow 2\text{AB}(\text{g})$ . What can we say about the mechanism of the reaction?
- if the reaction is fast it occurs in a single step
  - if the reaction is fast it occurs in multiple steps
  - if the reaction is slow it occurs in multiple steps

# Modern Chemistry Chapter 7 Test Answer Key

**Camilla Rothe**



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