

Vol. of 1 mol of ideal gas at STP/NTP =  $22.4 \text{ L}$

उत्स रान एरिा  
- 3 same sabse kam significant figure hoga pt. ke bad utne hi hoge ans me.

\* Significant fig  
• 0 to the left of the 1st non 0 digit are sig. not sig.  
eg 0.0012  $\Rightarrow$  2  
• Terminal 0 are not sig. if there is no decimal pt.  
eg 1000  $\Rightarrow$  1 fig.  
• 0 at end is sig. if there is decimal pt.  
• Exponential term are not considered sig.

अंश - 3/131  
axb=c  
- add/subst  
is most animal  
3 same sabse kam fig hoga vo sig. fig ans me hoga

\* LAW OF MASS CONSV  
• In a react<sup>n</sup>  $\rightarrow$  actual mass of reactant = mass of prod.  
Rak me zina use hua.

\* Except<sup>n</sup>  $\rightarrow$  Nuclear Rxn (E=mc<sup>2</sup>)  
• Pppwed<sup>n</sup>  $\rightarrow$  Balanced Rxn

# MOLE CONCEPT

\* LAW OF CONST PROPORT<sup>n</sup>  
• Elements in a comp are in fixed ratio by mass.  
eg CO<sub>2</sub>  
12 32g  $\Rightarrow$  3:8  
% of C =  $\frac{3}{11} \times 100$

Q. cal. the mass of C req to form 132g of CO<sub>2</sub>.  
 $\Rightarrow$  C + 2O  $\rightarrow$  CO<sub>2</sub>  
12g 32g 44g  
44g CO<sub>2</sub> ke lie 12g C  
 $\therefore$  132g ke lie  
=  $\frac{12}{44} \times 132 = 36\text{g}$

\* Law of Multiple Prop<sup>n</sup>  
• elements should be same or same fixed mass  
eg CO, CO<sub>2</sub> 1:2 ratio  
Other eg. N<sub>2</sub>O, N<sub>2</sub>O<sub>3</sub>, N<sub>2</sub>O<sub>5</sub>  
except<sup>n</sup>  $\rightarrow$  N<sub>2</sub>O<sub>3</sub>, N<sub>2</sub>O<sub>5</sub>

\* Avogadro's Law  
• Same (P, T, R, V) of molecules of gases same (mass eg! not of molar)  
X<sub>2</sub>, X<sub>3</sub>, X<sub>4</sub>  
R : R : R molecules  
2 : 3 : 4 atoms [different]

\* Gay Lussac's Law  
CONDINS: +T, P same  
• at least 2 gases.  
H<sub>2</sub> + Cl<sub>2</sub>  $\rightarrow$  2HCl  
1 : 1 : 2  
• Vol. K. Lie

\* DALTON'S ATOMIC THEORY  
• Atoms  $\rightarrow$  Neither creat nor destroyed

\* MOLE CONCEPT  
Atomic No. (P)  $\rightarrow$  mass no. (P+N)  
Mass of Proton =  $1.67 \times 10^{-27} \text{ kg}$   
Neutron =  $1.67 \times 10^{-27} \text{ kg}$   
Electron =  $9.1 \times 10^{-31} \text{ kg}$   
MP  $\approx$  1837 m<sub>e</sub>

\* Atomic mass unit (amu)  
1 amu =  $\frac{1}{12}$  Mass of <sup>12</sup>C  
1 amu =  $1.67 \times 10^{-27} \text{ kg}$   
eg Mass of <sup>12</sup>C = 12 amu  
Relative atomic mass  
• Unitless  
eg. mass. A.M. of C = 12

\* Molecular Mass  
• H<sub>2</sub>SO<sub>4</sub> = 98 amu  
 $\rightarrow$  Mass of 1 molecule of H<sub>2</sub>SO<sub>4</sub>  
FORMULA MASS: for ionic comp.

\* MOLE  
Amount of substance.  
• 1 mole = No particles (6.022 x 10<sup>23</sup>)  
1 amu  $\leq$  1.67 x 10<sup>-27</sup> kg  
Na = 6.022 x 10<sup>23</sup>  
1 amu y Na  $\leq$  1g

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**Victor M. Corman**



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