

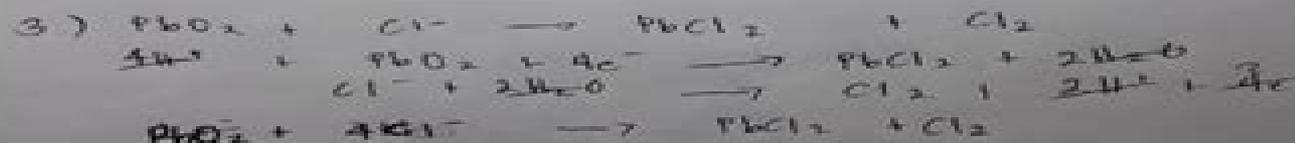
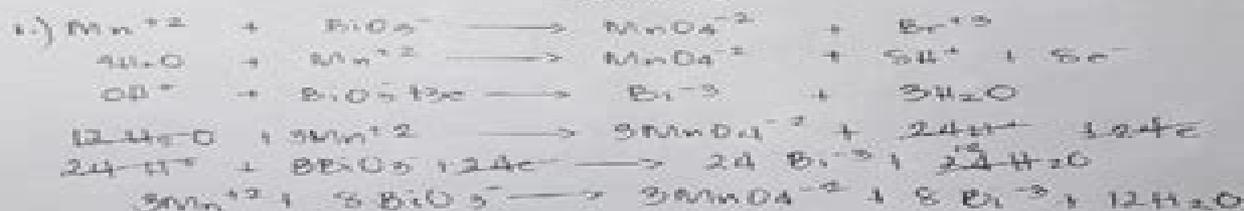
OXIDATION - REDUCTION REACTION

A. Balance the following ionic equations occurring in acidic solution.

1. $\text{Mn}^{+2} + \text{BiO}_3^- \rightarrow \text{MnO}_4^{-2} + \text{Bi}^{+3}$
2. $\text{NiO}_2 + \text{Mn}(\text{OH})_2 \rightarrow \text{Mn}_2\text{O}_3 + \text{Ni}(\text{OH})_2$
3. $\text{PbO}_2 + \text{Cl}^- \rightarrow \text{PbCl}_2 + \text{Cl}_2$
4. $\text{H}_2\text{PO}_3 + \text{Cr}_2\text{O}_7^{-2} \rightarrow \text{H}_2\text{PO}_4 + \text{Cr}^{+3}$
5. $\text{AsO}_3^{-2} + \text{Br}_2 \rightarrow \text{AsO}_4^{-2} + \text{Br}^-$

Oxidation - Reduction Reaction

A. Balance the following ionic equations occurring in acidic solution



Oxidation And Reduction Practice With Answers

Christian Drosten



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