



**AZEEM**

# PAKISTAN STUDIES

Intermediate Part-II

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# Pakistan Studies M A Part I

**JR Anderson**



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 $\text{H}_2\text{O}(\text{g}) \rightarrow \frac{1}{2} \text{O}_2(\text{g}) + \text{H}_2(\text{g})$   $\Delta h = 71.2 \text{ kJ}$   
 $\text{C}_2\text{H}_6(\text{g}) + \frac{7}{2} \text{O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 3 \text{ H}_2\text{O}(\text{g})$   $\Delta h = 283 \text{ kJ}$

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 $\text{C}_2\text{H}_4(\text{g}) + 3 \text{ O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 2 \text{ H}_2\text{O}(\text{l})$   $\Delta h = 1411 \text{ kJ/mole}$   
 $\text{C}_2\text{H}_6(\text{g}) + 7/2 \text{ O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 3 \text{ H}_2\text{O}(\text{l})$   $\Delta h = 1560 \text{ kJ/mole}$   
 $\text{H}_2(\text{g}) + 1/2 \text{ O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l})$   $\Delta h = 285 \text{ kJ/mole}$

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web thus for reaction 9 2 in which 1 mole of  $h_3o$  and 1 mole of  $oh$  combine to form 2 moles of  $h_2o$  h 55 8 kj as shown note that the enthalpy change  $h_{products} - h_{reactants}$  is positive if heat is absorbed that is if  $h_{products}$

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