

Topic: COLLIGATIVE PROPERTIES OF NONELECTROLYTES AND ELECTROLYTE SOLUTIONS

Colligative properties – are properties of a solution that depend only on the number and not on the identity of the solute particles.

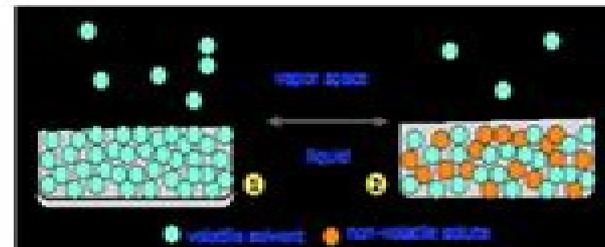
- vapor pressure lowering
- boiling point elevation and
- freezing point depression

Effect of solute concentration on the colligative properties of solutions

The concentration or amount of nonvolatile solute (i.e., a solute that does not have a vapor pressure of its own) in the solution has an effect on the colligative properties of solutions. The effect would depend on the ratio of the number of particles of solute and solvent in the solution and not on the identity of the solute. However, it is necessary to take into account whether the solute is an electrolyte or a nonelectrolyte.

Effects of electrolyte and nonelectrolyte on colligative properties solutions.

1. Vapor Pressure Lowering – When a nonvolatile solute is added to a pure solvent, the vapor pressure of the solution is lower than that of the pure solvent. This is because the solute particles interfere with the solvent molecules' ability to escape the liquid phase and enter the vapor phase. The more solute particles present, the greater the interference and the lower the vapor pressure.



2. Boiling point Elevation – The addition of a nonvolatile solute lowers the vapor pressure of the solution; consequently the temperature must be raised to restore the vapor pressure of the solution to the value conforming to the pure solvent. Specifically, the temperature at which the vapor pressure is 1 atm will be higher than the normal boiling point by an amount known as the boiling point elevation.

The lowering of the vapor pressure in a solution causes the boiling point of the solution to be higher than pure solvent



Boiling point elevation is the increase in the boiling point of a liquid when a non-volatile solute is added to it.

Properties Of Solutions Electrolytes And Nonelectrolytes Lab Report

Lauren Gardner



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Table of Contents Properties Of Solutions Electrolytes And Nonelectrolytes Lab Report

1. Understanding the eBook Properties Of Solutions Electrolytes And Nonelectrolytes Lab Report
 - The Rise of Digital Reading Properties Of Solutions Electrolytes And Nonelectrolytes Lab Report
 - Advantages of eBooks Over Traditional Books
2. Identifying Properties Of Solutions Electrolytes And Nonelectrolytes Lab Report
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Properties Of Solutions Electrolytes And Nonelectrolytes Lab Report

- User-Friendly Interface
- 4. Exploring eBook Recommendations from Properties Of Solutions Electrolytes And Nonelectrolytes Lab Report
 - Personalized Recommendations
 - Properties Of Solutions Electrolytes And Nonelectrolytes Lab Report User Reviews and Ratings
 - Properties Of Solutions Electrolytes And Nonelectrolytes Lab Report and Bestseller Lists
- 5. Accessing Properties Of Solutions Electrolytes And Nonelectrolytes Lab Report Free and Paid eBooks
 - Properties Of Solutions Electrolytes And Nonelectrolytes Lab Report Public Domain eBooks
 - Properties Of Solutions Electrolytes And Nonelectrolytes Lab Report eBook Subscription Services
 - Properties Of Solutions Electrolytes And Nonelectrolytes Lab Report Budget-Friendly Options
- 6. Navigating Properties Of Solutions Electrolytes And Nonelectrolytes Lab Report eBook Formats
 - ePub, PDF, MOBI, and More
 - Properties Of Solutions Electrolytes And Nonelectrolytes Lab Report Compatibility with Devices
 - Properties Of Solutions Electrolytes And Nonelectrolytes Lab Report Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Properties Of Solutions Electrolytes And Nonelectrolytes Lab Report
 - Highlighting and Note-Taking Properties Of Solutions Electrolytes And Nonelectrolytes Lab Report
 - Interactive Elements Properties Of Solutions Electrolytes And Nonelectrolytes Lab Report
- 8. Staying Engaged with Properties Of Solutions Electrolytes And Nonelectrolytes Lab Report
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Properties Of Solutions Electrolytes And Nonelectrolytes Lab Report
- 9. Balancing eBooks and Physical Books Properties Of Solutions Electrolytes And Nonelectrolytes Lab Report
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Properties Of Solutions Electrolytes And Nonelectrolytes Lab Report
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Properties Of Solutions Electrolytes And Nonelectrolytes Lab Report
 - Setting Reading Goals Properties Of Solutions Electrolytes And Nonelectrolytes Lab Report

- Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Properties Of Solutions Electrolytes And Nonelectrolytes Lab Report
 - Fact-Checking eBook Content of Properties Of Solutions Electrolytes And Nonelectrolytes Lab Report
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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