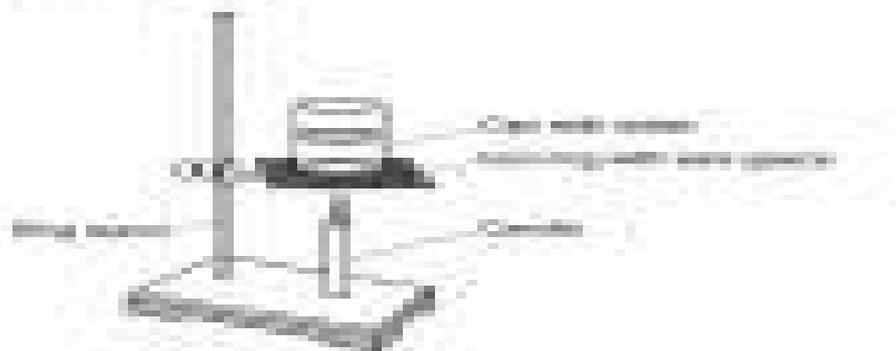


Part 1: Prepare a solution of approximately 100 mL of approximately 0.1 M sodium acetate in 200 mL of water.

Part 2: Prepare a solution of approximately 100 mL of approximately 0.1 M sodium acetate in 200 mL of water. Add a small amount of sodium acetate to the solution and stir. Record the pH of the solution.



Part 3: Prepare a solution of approximately 100 mL of approximately 0.1 M sodium acetate in 200 mL of water. Add a small amount of sodium acetate to the solution and stir. Record the pH of the solution. The pH of the solution is approximately 4.75. The pH of the solution is approximately 4.75. The pH of the solution is approximately 4.75.

Table 1: Titration Results

| Volume of NaOH added (mL) | pH | Approximate Concentration of Acetate (M) | Approximate Concentration of Acetic Acid (M) |
|---------------------------|------|--|--|
| 0 | 4.75 | 0.1 | 0 |

Part 4: Prepare a solution of approximately 100 mL of approximately 0.1 M sodium acetate in 200 mL of water. Add a small amount of sodium acetate to the solution and stir. Record the pH of the solution.

Part 5: Prepare a solution of approximately 100 mL of approximately 0.1 M sodium acetate in 200 mL of water. Add a small amount of sodium acetate to the solution and stir. Record the pH of the solution.

Part 6: Prepare a solution of approximately 100 mL of approximately 0.1 M sodium acetate in 200 mL of water. Add a small amount of sodium acetate to the solution and stir. Record the pH of the solution.

Heat Combustion Candle Lab Answers

DP Hallahan



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