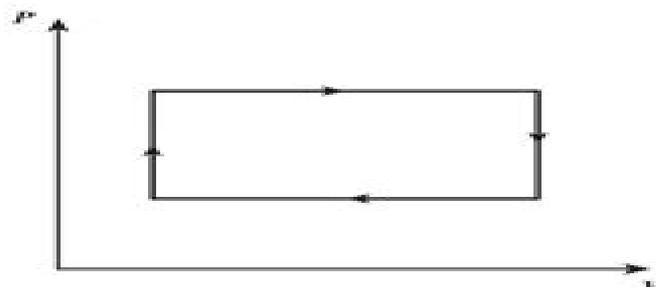


CHAPTER 9 PRACTICE QUESTIONS

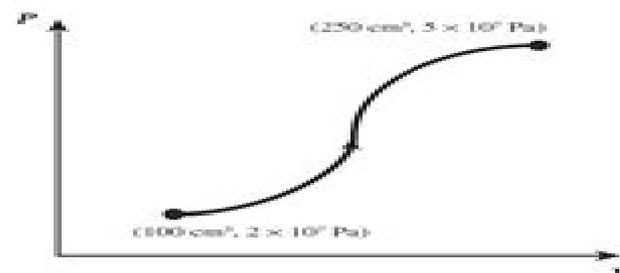
Directions: Complete the following problems as specified by each question, and then check your work using the solutions that follow. For extended, step-by-step solutions, access your Student Tools online.

- Water has a specific heat of $4186 \text{ J/kg}\cdot\text{K}$. How much heat would it take to raise the temperature of 100 g of water from 25°C to 75°C ?
- To change phase, a sample of a substance requires an amount of heat per unit mass known as the **latent heat** for that transformation. For evaporation, water has a latent heat of 2257 kJ/kg . How much heat is required to evaporate 30 g of water?
- How much heat is required to completely evaporate 50 g of water at an initial temperature of 50°C ? Refer to the previous two problems for relevant constants.
- If 3 mol of an atomic gas changes from a state $(0.1 \text{ m}^3, 1.0 \times 10^5 \text{ Pa})$ to $(0.2 \text{ m}^3, 3.0 \times 10^5 \text{ Pa})$ due to some arbitrary process, and absorbs 45 kJ of heat in the process, how much work is done? Is this an example of an engine?

- Does the gas in the following cycle lose or gain heat?



- What is the change in internal energy for a sample of molecular hydrogen undergoing the following process? Assume that the hydrogen atoms are connected rigidly.



High School Physics Questions And Answers

JS Bruner



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