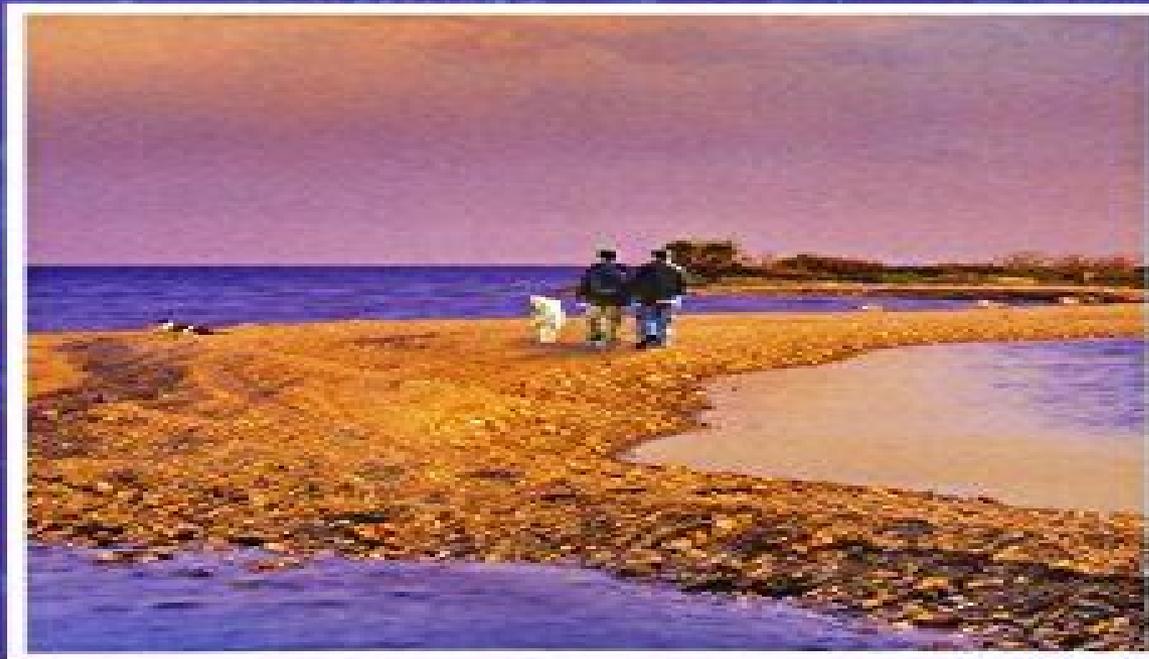




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web solution $v_1 = 2.50 \text{ l}$ $v_2 = 1.273 \text{ l}$ $t_1 = 303 \text{ k}$ $t_2 = 273 \text{ k}$ we will be applying charles law as pressure is constant here $v_1 t_1 = v_2 t_2$ or $v_2 = v_1 t_1 / t_2 = 2.50 \text{ l} \cdot 303 \text{ k} / 273 \text{ k} = 2.83 \text{ l}$ at 90 c a helium sample has a volume of 500 ml determine the temperature at which the volume of the liquid will become 240 ml