

## Thermochemistry Test Preview

### Matching

Match each item with the correct statement below.

- |                |                  |
|----------------|------------------|
| a. calorimeter | d. enthalpy      |
| b. calorie     | e. specific heat |
| c. joule       | f. heat capacity |

- \_\_\_\_\_ 1. quantity of heat needed to raise the temperature of 1 g of water by 1°C
- \_\_\_\_\_ 2. SI unit of energy
- \_\_\_\_\_ 3. quantity of heat needed to change the temperature of 1 g of a substance by 1°C
- \_\_\_\_\_ 4. quantity of heat needed to change the temperature of an object by 1°C
- \_\_\_\_\_ 5. device used to measure the heat absorbed or released during a chemical or physical process
- \_\_\_\_\_ 6. heat content of a system at constant pressure

Match each item with the correct statement below.

- |                                 |                     |
|---------------------------------|---------------------|
| a. heat of reaction             | d. heat of fusion   |
| b. heat of formation            | e. heat of solution |
| c. Hess's law of heat summation |                     |

- \_\_\_\_\_ 7. the enthalpy change for a chemical reaction exactly as it is written
- \_\_\_\_\_ 8. the enthalpy change caused by dissolving a substance
- \_\_\_\_\_ 9. the energy required to melt a solid at its melting point
- \_\_\_\_\_ 10. the change in enthalpy that accompanies the formation of a compound from its elements
- \_\_\_\_\_ 11. states that if you add two or more thermochemical equations to give a final equation, you can also add the heats of reaction to give the final heat of reaction

### Multiple Choice

Identify the letter of the choice that best completes the statement or answers the question.

- \_\_\_\_\_ 12. What happens to the energy produced by burning gasoline in a car engine?
  - a. The energy is lost as heat in the exhaust.
  - b. The energy is transformed into work to move the car.
  - c. The energy heats the parts of the engine.
  - d. all of the above
- \_\_\_\_\_ 13. A piece of metal is heated, then submerged in cool water. Which statement below describes what happens?
  - a. The temperature of the metal will increase.
  - b. The temperature of the water will increase.
  - c. The temperature of the water will decrease.
  - d. The temperature of the water will increase and the temperature of the metal will decrease.
- \_\_\_\_\_ 14. How does a calorie compare to a joule?
  - a. A calorie is smaller than a joule.
  - b. A calorie is larger than a joule.
  - c. A calorie is equal to a joule.
  - d. The relationship cannot be determined.
- \_\_\_\_\_ 15. What would likely happen if you were to touch the flask in which an endothermic reaction were occurring?
  - a. The flask would probably feel cooler than before the reaction started.
  - b. The flask would probably feel warmer than before the reaction started.
  - c. The flask would feel the same as before the reaction started.
  - d. none of the above

# Thermochemistry Practice Test A Answers

**SB Merriam**



**Thermochemistry Practice Test A Answers:**

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