

Physics

(www.tiwariacademy.in)
(Chapter 12)(Atoms)
(Class 12)
Exercises

Question 12.1:

Choose the correct alternative from the clues given at the end of the each statement:

- (a) The size of the atom in Thomson's model is _____ the atomic size in Rutherford's model. (much greater than/no different from/much less than.)
(b) In the ground state of _____ electrons are in stable equilibrium, while in _____ electrons always experience a net force. [Thomson's model/ Rutherford's model.]
(c) A classical atom based on _____ is doomed to collapse. [Thomson's model/ Rutherford's model.]
(d) An atom has a nearly continuous mass distribution in a _____ but has a highly non-uniform mass distribution in _____. [Thomson's model/ Rutherford's model.]
(e) The positively charged part of the atom possesses most of the mass in _____. [Rutherford's model/both the models.]

Answer 12.1:

- (a) The sizes of the atoms taken in Thomson's model and Rutherford's model have the same order of magnitude.
(b) In the ground state of Thomson's model, the electrons are in stable equilibrium. However, in Rutherford's model, the electrons always experience a net force.
(c) A classical atom based on Rutherford's model is doomed to collapse.
(d) An atom has a nearly continuous mass distribution in Thomson's model, but has a highly non-uniform mass distribution in Rutherford's model.
(e) The positively charged part of the atom possesses most of the mass in both the models.

Question 12.2:

Suppose you are given a chance to repeat the alpha-particle scattering experiment using a thin sheet of solid hydrogen in place of the gold foil. (Hydrogen is a solid at temperatures below 14 K.) What results do you expect?

Answer 12.2:

In the alpha-particle scattering experiment, if a thin sheet of solid hydrogen is used in place of a gold foil, then the scattering angle would not be large enough. This is because the mass of hydrogen (1.67×10^{-27} kg) is less than the mass of incident α -particles (6.64×10^{-27} kg). Thus, the mass of the scattering particle is more than the target nucleus (hydrogen). As a result, the α -particles would not bounce back if solid hydrogen is used in the α -particle scattering experiment.

Question 12.3:

What is the shortest wavelength present in the Paschen series of spectral lines?

Answer 12.3:

Rydberg's formula is given as:

$$\frac{hc}{\lambda} = 21.76 \times 10^{-19} \left[\frac{1}{n_1^2} - \frac{1}{n_2^2} \right]$$

Where,

h = Planck's constant = 6.6×10^{-34} Js and c = Speed of light = 3×10^8 m/s. (n_1 and n_2 are integers)

The shortest wavelength present in the Paschen series of the spectral lines is given for values $n_1 = 3$ and $n_2 = \infty$.

$$\frac{hc}{\lambda} = 21.76 \times 10^{-19} \left[\frac{1}{(3)^2} - \frac{1}{(\infty)^2} \right]$$

$$\begin{aligned} \lambda &= \frac{6.6 \times 10^{-34} \times 3 \times 10^8 \times 9}{21.76 \times 10^{-19}} \\ &= 8.189 \times 10^{-7} \text{ m} \\ &= 818.9 \text{ nm} \end{aligned}$$

Question 12.4:

A difference of 2.3 eV separates two energy levels in an atom. What is the frequency of radiation emitted when the atom makes a transition from the upper level to the lower level?

Physics Chapter 5 Review Answers

JR Anderson



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