

15.4 DETERMINING ENTHALPIES OF REACTION OR FUSION/BOILING

A third method used to determine enthalpies of reaction is the standard enthalpy of formation.

Others are calorimetry and Hess's law.

For example, the formation equation and the standard enthalpy of formation of carbon dioxide are:

$$\text{C (s)} + \text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g}) \quad \Delta H_f^\circ = -393.5 \text{ kJ/mol}$$

Steps for writing Formation Equations:

1. Write one mole of product in the state that has been specified
2. Write the reactant elements in their standard states
3. Choose reaction coefficients for the reactants to give a balanced equation yielding one mole of product.

Example 1:

Write the equation for formation of liquid ethanol

1. Start with one mole of product:



2. Write the reactant elements in their standard states



3. Balance the equation to yield one mole of product.



Example 2:

What is the formation equation for liquid carbonic acid?



Solution:



Example 3:



ΔH_f° is always for 1 mole of the product being formed

In general $\Delta H_f^\circ = (\sum \Delta H_f^\circ \text{ compounds formed}) - (\sum \Delta H_f^\circ \text{ element}_1 + \dots + \Delta H_f^\circ \text{ element}_n)$

ENTHALPY OF REACTION: ΔH_r°

is equal to the difference between the sum of enthalpies of the formation of products and that of the reactants.

$\Delta H_r^\circ = \text{enthalpy of Product} - \text{enthalpy of Reactants}$

$$\Delta H_r^\circ = \sum \Delta H_f^\circ \text{ Products} - \sum \Delta H_f^\circ \text{ Reactants}$$



$$\Delta H_r^\circ = [2\Delta H_f^\circ(\text{C}) + 2\Delta H_f^\circ(\text{D})] - [2\Delta H_f^\circ(\text{A}) + 2\Delta H_f^\circ(\text{B})]$$

Try the examples.

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Frank Jenkins



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1. Understanding the eBook Nelson Chemistry 12 Chapter 1 Solutions Webinn
 - The Rise of Digital Reading Nelson Chemistry 12 Chapter 1 Solutions Webinn
 - Advantages of eBooks Over Traditional Books
2. Identifying Nelson Chemistry 12 Chapter 1 Solutions Webinn
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Nelson Chemistry 12 Chapter 1 Solutions Webinn
 - User-Friendly Interface
4. Exploring eBook Recommendations from Nelson Chemistry 12 Chapter 1 Solutions Webinn
 - Personalized Recommendations
 - Nelson Chemistry 12 Chapter 1 Solutions Webinn User Reviews and Ratings
 - Nelson Chemistry 12 Chapter 1 Solutions Webinn and Bestseller Lists

5. Accessing Nelson Chemistry 12 Chapter 1 Solutions Webinn Free and Paid eBooks
 - Nelson Chemistry 12 Chapter 1 Solutions Webinn Public Domain eBooks
 - Nelson Chemistry 12 Chapter 1 Solutions Webinn eBook Subscription Services
 - Nelson Chemistry 12 Chapter 1 Solutions Webinn Budget-Friendly Options
6. Navigating Nelson Chemistry 12 Chapter 1 Solutions Webinn eBook Formats
 - ePub, PDF, MOBI, and More
 - Nelson Chemistry 12 Chapter 1 Solutions Webinn Compatibility with Devices
 - Nelson Chemistry 12 Chapter 1 Solutions Webinn Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Nelson Chemistry 12 Chapter 1 Solutions Webinn
 - Highlighting and Note-Taking Nelson Chemistry 12 Chapter 1 Solutions Webinn
 - Interactive Elements Nelson Chemistry 12 Chapter 1 Solutions Webinn
8. Staying Engaged with Nelson Chemistry 12 Chapter 1 Solutions Webinn
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Nelson Chemistry 12 Chapter 1 Solutions Webinn
9. Balancing eBooks and Physical Books Nelson Chemistry 12 Chapter 1 Solutions Webinn
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Nelson Chemistry 12 Chapter 1 Solutions Webinn
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Nelson Chemistry 12 Chapter 1 Solutions Webinn
 - Setting Reading Goals Nelson Chemistry 12 Chapter 1 Solutions Webinn
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Nelson Chemistry 12 Chapter 1 Solutions Webinn
 - Fact-Checking eBook Content of Nelson Chemistry 12 Chapter 1 Solutions Webinn
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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