

# Basic Concepts

## Section 1.2 Solutions

1.1.1 If the current in an electric conductor is 3.4 A, how many coulombs of charge pass any point in a 30 s interval?

**Solution:**

$$\begin{aligned} I &= 3.4 \text{ A}, \Delta t = 30 \text{ s} \\ Q &= I \Delta t \\ Q &= 102 \text{ C} \end{aligned}$$

1.1.2 Determine the time interval required for a 0.3 A battery charger to deliver 4800 C.

**Solution:**

$$\begin{aligned} I &= 0.3 \text{ A}, Q = 4800 \text{ C} \\ \Delta t &= \frac{Q}{I} \\ \Delta t &= 1600 \text{ s} \end{aligned}$$

1.1.3 If a 90 V battery delivers 480 C in 3 s, find (a) the amount of charge delivered and (b) the current produced.

**Solution:**

$$\begin{aligned} V &= 90 \text{ V}, \Delta Q = 480 \text{ C in } 3 \text{ s} \\ \text{a. } \Delta Q &= 480 \text{ C} \\ \Delta Q &= 480 \text{ C} \\ \text{b. } I &= \frac{\Delta Q}{\Delta t}, \Delta t = 3 \text{ s} \\ I &= 160 \text{ A} \end{aligned}$$

1.1.4 The current in a conductor is 1.7 A. How many coulombs of charge pass any point in a time interval of 3.3 min?

**Solution:**

$$\begin{aligned} I &= 1.7 \text{ A}, \Delta t = 3.3 \text{ min} = 198 \text{ s} \\ Q &= I \Delta t \\ Q &= 336 \text{ C} \end{aligned}$$

# Chapter 2 Solution University Of Rochester

**Jin-Ying Zhang**



## Chapter 2 Solution University Of Rochester:

**Water and Aqueous Solutions at Subzero Temperatures** Felix Franks, 2013-11-11 This Volume the last of the series is devoted to water in its metastable forms especially at sub zero temperatures The past few years have witnessed an increasing interest in supercooled water and amorphous ice If the properties of liquid water in the normal temperature range are already eccentric then they become exceedingly so below the normal freezing point in the metastable temperature range Water can be supercooled to 39 C without too much effort and most of its physical properties show a remarkable temperature dependence under these conditions Although adequate explanations are still lacking the time has come to review available knowledge The study of amorphous ice that is the solid formed when water vapor is condensed on a very cold surface is of longer standing It has achieved renewed interest because it may serve as a model for the liquid state There is currently a debate whether or not a close structural relationship exists between amorphous ice and supercooled water The nucleation and growth of ice in supercooled water and aqueous solutions is also still one of those grey areas of research although these topics have received considerable attention from chemists and physicists over the past two decades Even now the relationships between degree of supercooling nucleation kinetics crystal growth kinetics cooling rate and solute concentration are somewhat obscure Nevertheless at the empirical level much progress has been made because these topics are of considerable importance to biologists technologists atmospheric physicists and glaciologists [Water, a Comprehensive Treatise: Water and aqueous solutions at subzero temperatures](#) Felix Franks, 1972 V 4 Aqueous solutions of amphiphiles and macromolecules Author subject and compound indexes *The Chemistry of Alkenes* Saul Patai, Jacob Zabicky, 1964 *Solutions of Population Balance Equations for Growing Tetrahymena Pyriformis Using a Successive Generations Approach* Jia-Jer Liou, 1996 *Advances in Terminal Care Research and Application: 2012 Edition*, 2012-12-26 *Advances in Terminal Care Research and Application 2012 Edition* is a ScholarlyBrief that delivers timely authoritative comprehensive and specialized information about Terminal Care in a concise format The editors have built *Advances in Terminal Care Research and Application 2012 Edition* on the vast information databases of ScholarlyNews You can expect the information about Terminal Care in this eBook to be deeper than what you can access anywhere else as well as consistently reliable authoritative informed and relevant The content of *Advances in Terminal Care Research and Application 2012 Edition* has been produced by the world's leading scientists engineers analysts research institutions and companies All of the content is from peer reviewed sources and all of it is written assembled and edited by the editors at ScholarlyEditions and available exclusively from us You now have a source you can cite with authority confidence and credibility More information is available at <http://www.ScholarlyEditions.com> **Quantitative Chemical Analysis** Daniel C. Harris, 2010-04-30 The most widely used analytical chemistry textbook in the world Dan Harris's *Quantitative Chemical Analysis* provides a sound physical understanding of the principles of analytical chemistry showing how these principles are

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