

Section 12.1

Section 12.1

Answer complete answer multiple-choice questions that are provided.

1. _____
 - a. _____
 - (a) correct, in terms of all four parts and products.
 - (b) correct, in terms of number of reactants and products.
 - (c) correct, in terms of work done on work done by system.
 - (d) correct, in terms of enthalpy change in the reaction.
 - b. _____
 - (a) the mass number of all the reactants combined is equal to the mass number of all the products combined.
 - (b) the mass of carbon produced is equal to the mass of carbon consumed.
 - (c) the number of moles of hydrogen that will react with a certain quantity of oxygen.
 - (d) the amount of energy required to react the reactants listed.
 - c. _____
 - (a) correct, in terms of energy change in the reaction.
 - (b) correct, in terms of the number of moles of reactants and products.
 - (c) correct, in terms of the number of moles of reactants and products.
 - (d) correct, in terms of the number of moles of reactants and products.
 - d. _____
 - (a) correct, in terms of the number of moles of reactants and products.
 - (b) correct, in terms of the number of moles of reactants and products.
 - (c) correct, in terms of the number of moles of reactants and products.
 - (d) correct, in terms of the number of moles of reactants and products.

2. Answer the two questions based on the following balanced equation:



a. Balance the equation:



b. What is the coefficient of H_2 in the equation?

c. How many moles of H_2 are required to react with 1 mole of O_2 to produce 2 moles of H_2O ?

d. How many moles of H_2 are required to react with 1 mole of O_2 to produce 2 moles of H_2O ?

e. How many moles of H_2 are required to react with 1 mole of O_2 to produce 2 moles of H_2O ?

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