

name Key

DUE date/TEST date 10/24 - Mon 10/25 - Tue

# Practice Problems Unit 5 - Nuclear Chemistry

Directions: Answer each of the following questions. Show ALL work in the space provided with units and correct sig figs to receive credit.

## Practice Problems - Nuclear Reactions

1. Fill out the following chart to help predict types of radiation.

Type of Radiation	How can you recognize that problem?	
beta decay	$\frac{0}{-1}e$	mass # ↓ 0 atomic # ↓ 1
alpha decay	$\frac{4}{2}He$	mass # ↓ 4 atomic # ↓ 2
gamma decay		γ - no change

2. Fill out the chart based on your knowledge of nuclear reactions.

Element	Protons	Neutrons	Electrons	Stable?	Type of Radiation	Balanced Nuclear Reaction
carbon-14	6	8	6	no	beta decay	$^{14}_6C \rightarrow ^{14}_7N + ^0_{-1}e$
radium-226	88	$226 - 88 = 138$	88	no	alpha decay	$^{226}_{88}Ra \rightarrow ^4_2He + ^{222}_{86}Rn$
$^{65}Cu$	29	$65 - 29 = 36$	29	no	beta decay	$^{65}_{29}Cu \rightarrow ^0_{-1}e + ^{65}_{30}Zn$
$^{127m}I$	53	$127 - 53 = 74$	53	no	gamma decay	$^{127m}_{53}I \rightarrow ^0_0\gamma + ^{127}_{53}I$
$^{186m}Re$	75	$186 - 75 = 111$	75	no	gamma decay	$^{186m}_{75}Re \rightarrow ^0_0\gamma + ^{186}_{75}Re$
nobelium-259	102	$259 - 102 = 157$	102	no	alpha decay	$^{259}_{102}No \rightarrow ^4_2He + ^{255}_{100}Fm$

3. How are the mass number and atomic number of the nucleus affected by the following:

- Beta decay mass # same atomic # ↑ 1
- Alpha decay mass # ↓ 4 atomic # ↓ 2
- Gamma decay mass # same atomic # same

4. Complete and balance each of the following nuclear reactions.

- $^{28}Al \rightarrow ^{28}Si + ^0_{-1}e$
- $^{210}_{83}Bi \rightarrow ^{206}_{81}Tl + ^4_2He$

5. Bromine -82 decays to produce Krypton - 82.

- What type of decay is this? justify your answer! beta decay - mass stayed the same + atomic # ↑ 1
- Write the nuclear reaction for this decay.  $^{82}_{35}Br \rightarrow ^{82}_{36}Kr + ^0_{-1}e$

6. Neptunium - 237 undergoes decay to Protactinium - 233.

- What type of decay is this? justify your answer! alpha decay → mass # ↓ 4 + atomic # ↓ 2
- Write the nuclear reaction for this decay.  $^{237}_{93}Np \rightarrow ^{233}_{91}Pa + ^4_2He$

7. Write a balanced nuclear reaction for the beta decay of:

- Nitrogen - 16  $^{16}_7N \rightarrow ^0_{-1}e + ^{16}_8O$
- Potassium - 40  $^{40}_{19}K \rightarrow ^0_{-1}e + ^{40}_{20}Ca$

8. Write a balanced nuclear reaction for the alpha decay of:

- Plutonium - 244  $^{244}_{94}Pu \rightarrow ^4_2He + ^{240}_{92}U$
- Strontium - 90  $^{90}_{38}Sr \rightarrow ^4_2He + ^{86}_{36}Kr$

# Chapter 22 Nuclear Chemistry Section 1 Review Answers

**Jicheng Xie**



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