

Chemistry I Chapter 5 Study Guide

Multiple Choice

Identify the choice that best completes the statement or answers the question.

- _____ 1. The noble gases, such as helium and xenon, are non-reactive because:
- they have completely filled outer electron shells.
 - they are chemically unstable.
 - they are unusually large atoms.
 - they have been around longest on the earth.
- _____ 2. Elements in the same column on the periodic table have similar chemical reactivity because:
- they contain similar amounts of energy.
 - they have the same number of valence shell electrons.
 - atoms prefer to have completely full or completely empty valence shells.
 - they have similar numbers of isotopes.
- _____ 3. Chlorine, element number 17, is located in the third row and the next-to-the-last group of the periodic table. How many electron shells would be completely filled by a neutral atom of chlorine? How many electrons would be left over?
- 1 shell filled, 7 electrons left over
 - 1 shell filled, 1 electron left over
 - 2 shells filled, 7 electrons left over
 - 2 shells filled, 1 electron left over
- _____ 4. How many energy levels are filled in a krypton atom?

36
Kr
Krypton
83.80

- Three.
 - Four.
 - Five.
 - Six.
- _____ 5. The electronegativity of an element is defined as:
- the attraction an atom has for the shared pair of electrons in a covalent bond.
 - the attraction an atom has for the shared pair of electrons in an ionic bond.
 - the amount of static electricity an element produces.
 - the amount of charge one atom of an element produces.
- _____ 6. Which of the following is true for valence electrons?
- Valence electrons are always located in the innermost energy level.
 - Valence electrons may be lost or gained by an atom when bonding occurs with another atom.
 - Each element in the periodic table has a different number of valence electrons.
 - Valence electrons are found only in radioactive isotopes.
- _____ 7. Mendeleev arranged the known chemical elements in a table according to increasing
- atomic number.
 - number of electrons.
 - number of protons.
 - mass.
- _____ 8. Which list of elements contains only metals?
- carbon, iodine, tin
 - tin, copper, cesium
 - helium, iron, copper
 - iodine, carbon, argon
- _____ 9. As you move from left to right across a period, the number of valence electrons
- increases.
 - stays the same.
 - increases and then decreases.
 - decreases.
- _____ 10. A member of the boron family has three valence electrons, while a member of the nitrogen family has
- none.
 - four.

Chapter 5 Chemistry Answers

John Giorgi



Chapter 5 Chemistry Answers:

Wastewater Operator Certification Study Guide John Giorgi,2011-01-12

CSIR NET Chemical Sciences Question

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Study Guide, Answer Book to Accompany Organic Chemistry Alan S. Wingrove,Robert L. Caret,1981

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American Illustrated Magazine, 1913 American Magazine, 1913 *The Independent*, 1914 **The Independent** William Livingston, 1913 General Elementary Science William Briggs, 1898 Chemistry for the Utterly Confused John T. Moore, Richard H. Langley, 2007-06-01 Banish bafflement in this tough subject. From formulas and lab techniques to the periodic table. *Chemistry for the Utterly Confused* focuses on the areas of maximum confusion and breaks down the most difficult chemistry topics into easy to understand concepts. This invaluable guide also teaches problem solving skills you need to master this imposing subject. Whether you're in high school, in college or simply brushing up on chemistry knowledge, this fun, easily accessible book will make understanding chemistry a breeze. Answer Book for Chemistry, Sixth Edition Charles E. Mortimer, 1986 The Chautauquan, 1885 *Complete Preparation for the MCAT* Williams & Wilkins Review, 1998-04 Here is the most respected test prep book for the Medical College Admission Test you can buy, featuring an active learning approach for a better understanding of the exam's content and a better chance for success. Unique to this guide are coverage of all recent changes in the MCAT plus a step by step plan for sharpening cognitive skills, developing problem solving skills and critical thinking. This thorough guide replaces expensive test preparation courses while giving students exactly what they need to get ready for the MCAT. *A Text-book of Inorganic Chemistry for University Students* James Riddick Partington, 1925 The Success Manual for General Chemistry Elizabeth Kean, Catherine Middlecamp, 1986

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