

CHAPTER 6

STUDY GUIDE

The Periodic Table and Periodic Law

Section 6.1 Development of the Modern Periodic Table

In your textbook, read about the history of the periodic table's development.

Use each of the terms below just once to complete the passage.

octaves	atomic mass	atomic number	size
elements	properties	Henry Mosely	eight
protons	periodic law	Dmitri Mendeleev	<u>accepted</u>

The table below was developed by John Newlands and is based on a relationship called the law of (1) properties. According to this law, the properties of the elements repeated every (2) eight elements. Thus, for example, element two and element (3) nine have similar properties. The law of octaves did not work for all the known elements and was not generally (4) accepted.

1	2	3	4	5	6	7
H	Li	G	Be	C	N	O
8	9	10	11	12	13	15
F	Na	Mg	Al	Si	P	S

The first periodic table is mostly credited to (5) Mendeleev. In his table, the elements were arranged according to increasing (6) atomic mass. One important result of this table was that the existence and properties of undiscovered (7) elements could be predicted.

The element in the modern periodic table are arranged according to increasing (8) atomic number, as a result of the work of (9) Mosely. This arrangement is based on number of (10) protons in the nucleus of an atom of the element. The modern form of the periodic table results in the (11) periodic law which states that when elements are arranged according to increasing atomic number, there is a periodic repetition of their chemical and physical (12) properties.

Chapter 6 Study Guide The Periodic Table And Law Answers

Peter P. Berlow



Chapter 6 Study Guide The Periodic Table And Law Answers:

Unveiling the Magic of Words: A Review of "**Chapter 6 Study Guide The Periodic Table And Law Answers**"

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Table of Contents Chapter 6 Study Guide The Periodic Table And Law Answers

1. Understanding the eBook Chapter 6 Study Guide The Periodic Table And Law Answers
 - The Rise of Digital Reading Chapter 6 Study Guide The Periodic Table And Law Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 6 Study Guide The Periodic Table And Law Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 6 Study Guide The Periodic Table And Law Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 6 Study Guide The Periodic Table And Law Answers
 - Personalized Recommendations
 - Chapter 6 Study Guide The Periodic Table And Law Answers User Reviews and Ratings
 - Chapter 6 Study Guide The Periodic Table And Law Answers and Bestseller Lists

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5. Accessing Chapter 6 Study Guide The Periodic Table And Law Answers Free and Paid eBooks
 - Chapter 6 Study Guide The Periodic Table And Law Answers Public Domain eBooks
 - Chapter 6 Study Guide The Periodic Table And Law Answers eBook Subscription Services
 - Chapter 6 Study Guide The Periodic Table And Law Answers Budget-Friendly Options
6. Navigating Chapter 6 Study Guide The Periodic Table And Law Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chapter 6 Study Guide The Periodic Table And Law Answers Compatibility with Devices
 - Chapter 6 Study Guide The Periodic Table And Law Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 6 Study Guide The Periodic Table And Law Answers
 - Highlighting and Note-Taking Chapter 6 Study Guide The Periodic Table And Law Answers
 - Interactive Elements Chapter 6 Study Guide The Periodic Table And Law Answers
8. Staying Engaged with Chapter 6 Study Guide The Periodic Table And Law Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 6 Study Guide The Periodic Table And Law Answers
9. Balancing eBooks and Physical Books Chapter 6 Study Guide The Periodic Table And Law Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 6 Study Guide The Periodic Table And Law Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chapter 6 Study Guide The Periodic Table And Law Answers
 - Setting Reading Goals Chapter 6 Study Guide The Periodic Table And Law Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chapter 6 Study Guide The Periodic Table And Law Answers
 - Fact-Checking eBook Content of Chapter 6 Study Guide The Periodic Table And Law Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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