

# Chapter 6

## Chapter 6: Thermochemistry

### 6.2 - Enthalpy and Calorimetry

Enthalpy

$$H = E + PV$$

For a chemical reaction, enthalpy change is given by the equation

$$\Delta H = H_{\text{products}} - H_{\text{reactants}}$$

Products of a reaction have a greater enthalpy ( $\Delta H$ ):

- $\Delta H$  is positive
- heat is absorbed by the system (endothermic reaction)

Reactants of a reaction have a greater enthalpy ( $\Delta H$ ):

- $\Delta H$  is negative
- heat is released by the system (exothermic reaction)

Calorimetry

Constant-pressure calorimetry

change in enthalpy = change in heat

$$\Delta H = q_p$$

# Chapter 6 Thermochemistry Faculty Rmu

**K Morrison**



## **Chapter 6 Thermochemistry Faculty Rmu:**

## **Chapter 6 Thermochemistry Faculty Rmu** Book Review: Unveiling the Magic of Language

In an electronic digital era where connections and knowledge reign supreme, the enchanting power of language has become apparent than ever. Its power to stir emotions, provoke thought, and instigate transformation is actually remarkable. This extraordinary book, aptly titled "**Chapter 6 Thermochemistry Faculty Rmu**," compiled by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we shall delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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