

Review for Ch. 8 Test

8.1 Describing Chemical Reactions

- know how to write word equations (i.e. you have be able to name compounds, including acids, ionic, and covalent compounds from Ch. 7)
- know how to write a formula equation
- What 2 things do coefficients represent in an equation? **Relative number of moles of a substance ; relative number of molecules needed for a reaction to occur**
- know common symbols used in a reaction
- know 4 evidences of a chemical reaction:
 1. **energy in the form of heat of light**
 2. **production of a gas**
 3. **change in color**
 4. **formation of a precipitate**
- know the Law of conservation of matter; **matter can neither be created nor destroyed in a chemical reaction (which means, mass of reactants must equal the mass of products)**
- know what a chemical equation represents: **relative amounts of reactants and products**
- know how to balance chemical equations (combustion reactions are particularly tricky, practice!)

1. Complete the equation and write a word equation for the following reaction: $\text{ZnO} + \text{FeSO}_4 \rightleftharpoons$



Zinc oxide reacts with iron (II) sulfate to yield zinc sulfate and iron (II) oxide

2. Write a formula equation for the reaction of a solution of AgCl with potassium metal:



3. What elements are always diatomic gases at room temperature?

All of the halogens (Group 17) , H_2 , N_2 , O_2

4. Can you balance these reactions:



Chapter 8 Chemical Equations And Reactions Test Answer Key

J. G. R. Briggs



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