



COVALENT BONDING

Practice Problems

Do your answers to the following problems.

SECTION 8.1 MOLECULAR COMPOUNDS

- Classify each of the following as either ionic or molecular.
 - NaCl
 - H_2O
 - HCl
 - KNO_3
 - H_2SO_4
 - H_2O_2
- Write the formula for the following molecular compounds.
 - nitrogen dioxide
 - carbon dioxide
 - carbon monoxide
 - nitrogen monoxide
 - nitrogen trioxide
 - nitrogen pentoxide

- Write the molecular formula for each of the following compounds.
 - nitrogen dioxide
 - carbon dioxide
 - carbon monoxide
 - nitrogen monoxide
 - nitrogen trioxide
 - nitrogen pentoxide
- Write the molecular formula for each of the following compounds.
 - nitrogen dioxide
 - carbon dioxide
 - carbon monoxide
 - nitrogen monoxide
 - nitrogen trioxide
 - nitrogen pentoxide

SECTION 8.2 THE NATURE OF COVALENT BONDING

- Draw the Lewis structure for hydrogen fluoride. $\text{H}-\text{F}$
- Draw the Lewis structure for carbon dioxide. $\text{O}=\text{C}=\text{O}$
- Draw the Lewis structure for methane. C with four H atoms bonded to it.
- Draw the Lewis structure for ethane. $\text{H}-\text{C}-\text{C}-\text{H}$ with appropriate H atoms.
- Draw the Lewis structure for acetylene. $\text{H}-\text{C}\equiv\text{C}-\text{H}$
- Draw the Lewis structure for hydrogen cyanide. $\text{H}-\text{C}\equiv\text{N}$



SECTION 8.3 BONDING: THE ORBITALS

- Which orbital and sub-orbital are represented by each of the following?
 - $2s$
 - $2p$
 - $3s$
 - $3p$
 - $4s$
 - $4p$
- Which orbital and sub-orbital are represented by each of the following?
 - $2s$
 - $2p$
 - $3s$
 - $3p$
 - $4s$
 - $4p$

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Wolfgang Guggemos



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