

CHAPTER 6 REVIEW*Chemical Bonding***MIXED REVIEW**

SHORT ANSWER Answer the following questions in the space provided.

1. Name the type of energy that is a measure of strength for each of the following types of bonds:

_____ **lattice energy** _____ a. ionic bond
 _____ **bond energy** _____ b. covalent bond
 _____ **enthalpy of vaporization** _____ c. metallic bond

2. Use the electronegativity values shown in **Figure 20**, on page 161 of the text, to determine whether each of the following bonds is nonpolar covalent, polar covalent, or ionic.

_____ **ionic** _____ a. H—F _____ **nonpolar covalent** _____ d. H—H
 _____ **ionic** _____ b. Na—Cl _____ **polar covalent** _____ e. H—C
 _____ **polar covalent** _____ c. H—O _____ **polar covalent** _____ f. H—N

3. How is a hydrogen bond different from an ionic or covalent bond?

A hydrogen bond is a dipole-dipole attraction between a partially positive hydrogen atom and the unshared electron pair of a strongly electronegative atom such as O, N, or F. Unlike ionic or covalent bonds, in which electrons are given up or shared, the hydrogen bond is a weaker attraction. Hydrogen bonds are generally intermolecular, while ionic and covalent bonds occur between ions or atoms respectively.

4. H₂S and H₂O have similar structures and their central atoms belong to the same group. Yet H₂S is a gas at room temperature and H₂O is a liquid. Use bonding principles to explain why this is.

Oxygen has higher electronegativity than sulfur, which creates a highly polar bond. Increased polarity in H₂O bonds means a stronger intermolecular attraction, making water a liquid at room temperature. Hydrogen bonding exists between water molecules, but not between hydrogen sulfide molecules.

Chapter 8 Covalent Bonding Test B Answers

**Alan Sherman, Sharon
Sherman, Leonard Russikoff**



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web there is a great virtual frog dissection lab online that is perfect for biology and anatomy students there are questions within the lab but no way to submit them to the instructor

student laboratory frog dissection external - Dec 29 2022

web 4 avoid contact with preservative chemicals keep your hands covered with the plastic gloves at all times rinse the frog completely before dissection 5 use the proper

[lab frog dissection introduction astephensscience](#) - Jul 24 2022

web a lab sheet for a 5 day dissection lab only three of the days are actually cutting open the frog as the other two are focused on providing students with the other parts of the