



# COVALENT BONDING

## Practice Problems

Do your answers to the following problems.

### SECTION 8.1 MOLECULAR COMPOUNDS

- Classify each of the following as either ionic or molecular.
  - $\text{NaCl}$
  - $\text{H}_2\text{O}$
  - $\text{HCl}$
  - $\text{KNO}_3$
  - $\text{C}_2\text{H}_6$
  - $\text{H}_2\text{SO}_4$
  - $\text{CaCl}_2$
  - $\text{C}_2\text{H}_5\text{OH}$
  - $\text{CO}_2$

- Classify each of the following as either ionic or molecular compound. Write the name of the compound.
- $\text{K}_2\text{SO}_4$
- $\text{C}_2\text{H}_6$
- $\text{H}_2\text{O}$
- $\text{H}_2\text{SO}_4$
- $\text{CaCl}_2$
- $\text{C}_2\text{H}_5\text{OH}$
- $\text{CO}_2$

### SECTION 8.2 THE NATURE OF COVALENT BONDING

- Show covalent bonding in hydrogen fluoride.  $\text{H}-\text{F}$
- Show covalent bonding in methane.  $\text{C}$  with four  $\text{H}$  atoms bonded to it.
- Show covalent bonding in carbon dioxide.  $\text{C}$  with two  $\text{O}$  atoms bonded to it.
- Show covalent bonding in ethane.  $\text{H}-\text{C}-\text{C}-\text{H}$
- Show covalent bonding in acetylene.  $\text{H}-\text{C}\equiv\text{C}-\text{H}$
- Draw Lewis structures for the following molecules.
  - $\text{H}_2\text{O}$
  - $\text{CO}_2$
  - $\text{C}_2\text{H}_6$
  - $\text{C}_2\text{H}_4$
  - $\text{C}_2\text{H}_2$
  - $\text{C}_2\text{H}_6\text{O}$
  - $\text{C}_2\text{H}_6\text{N}_2$
  - $\text{C}_2\text{H}_6\text{N}_2\text{O}$
  - $\text{C}_2\text{H}_6\text{N}_2\text{O}_2$
  - $\text{C}_2\text{H}_6\text{N}_2\text{O}_3$

### SECTION 8.3 BONDING: THE ORBITALS

- Which orbital and sub-orbital are represented by  $\text{sp}^3$ ? The diagram shows the four orbitals.
- Which orbital and sub-orbital form the  $\text{sp}^2$  hybrid orbitals shown?

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**Ying Liu**



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