



COVALENT BONDING

Practice Problems

Do your answers under the following problems.

SECTION 8.1 MOLECULAR COMPOUNDS

- Classify each of the following as either ionic or molecular.
 - NaCl
 - H_2O
 - HCl
 - KNO_3
 - H_2SO_4
 - H_2O_2
 - C_2H_6
 - CO_2

- Classify each of the following as either ionic or molecular compound. Write the formula for each compound and name the compound.
 - Two elements, one a metal and one a nonmetal, combine to form an ionic compound. The compound contains two metal ions for every one nonmetal ion.
 - Two elements, one a metal and one a nonmetal, combine to form an ionic compound. The compound contains two nonmetal ions for every one metal ion.

SECTION 8.2 THE NATURE OF COVALENT BONDING

- Show covalent bonding in hydrogen fluoride. $\text{H}-\text{F}$
- Show covalent bonding in methane. CH_4
- Show covalent bonding in carbon dioxide. CO_2
- Show covalent bonding in acetylene. $\text{H}-\text{C}\equiv\text{C}-\text{H}$
- Draw Lewis structures for each of the following.
 - H_2O
 - H_2O_2
 - C_2H_6
 - C_2H_4
 - C_2H_2
 - $\text{C}_2\text{H}_6\text{O}$
 - $\text{C}_2\text{H}_6\text{N}_2$
 - $\text{C}_2\text{H}_4\text{N}_2$
 - $\text{C}_2\text{H}_4\text{O}_2$
 - $\text{C}_2\text{H}_4\text{O}$
 - $\text{C}_2\text{H}_2\text{O}_2$
 - $\text{C}_2\text{H}_2\text{O}$
 - $\text{C}_2\text{H}_2\text{N}_2$
 - $\text{C}_2\text{H}_2\text{N}$
 - $\text{C}_2\text{H}_2\text{O}_3$
 - $\text{C}_2\text{H}_2\text{O}_4$
 - $\text{C}_2\text{H}_2\text{O}_5$
 - $\text{C}_2\text{H}_2\text{O}_6$
 - $\text{C}_2\text{H}_2\text{O}_7$
 - $\text{C}_2\text{H}_2\text{O}_8$
 - $\text{C}_2\text{H}_2\text{O}_9$
 - $\text{C}_2\text{H}_2\text{O}_{10}$
 - $\text{C}_2\text{H}_2\text{O}_{11}$
 - $\text{C}_2\text{H}_2\text{O}_{12}$
 - $\text{C}_2\text{H}_2\text{O}_{13}$
 - $\text{C}_2\text{H}_2\text{O}_{14}$
 - $\text{C}_2\text{H}_2\text{O}_{15}$
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 - $\text{C}_2\text{H}_2\text{O}_{46}$
 - $\text{C}_2\text{H}_2\text{O}_{47}$
 - $\text{C}_2\text{H}_2\text{O}_{48}$
 - $\text{C}_2\text{H}_2\text{O}_{49}$
 - $\text{C}_2\text{H}_2\text{O}_{50}$

SECTION 8.3 BONDING: THE ORBITALS

- Which orbital and sub-orbital do you combine to form the sigma bond in H_2 ? The sigma bond in H_2 is formed from which orbitals?
- Which orbital and sub-orbital do you combine to form the sigma bond in H_2 ? The sigma bond in H_2 is formed from which orbitals?

Chapter 8 Covalent Bonding Work Answers Prentice Hall

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