

## Prentice Hall Physical Science Chapter 4 Atomic Structure

### 4.1 Studying Atoms

- A. **Ancient Greek Models of Atoms**
  - Democritus believed all matter was made of extremely small particles that could not be divided
  - he called the particles atoms (from the Greek atomos meaning indivisible)
  - he thought there were different types of atoms with different properties (liquids had round, smooth atoms and solids were rough and prickly)
- B. **Dalton's Atomic Theory**
  - he proved atoms exist by discovering that compounds have a fixed composition
  - Dalton's Theory:
    - all elements are composed of atoms
    - all atoms of the same element have the same mass, and atoms of different elements have different masses
    - compounds contain atoms of more than one element
    - in a particular compound, atoms of different element always combine in the same way
  - he pictured atoms as solid spheres
  - in time it was found that not all of this was completely correct
- C. **Thomson's Model of the Atom**
  - some objects have a positive or negative charge and objects with the same charge repel and objects with opposite charges attract
  - Thomson used an electric current to learn more about atoms
  - he passed an electric current through a gas and a glowing beam appeared
  - the beam was repelled by a negatively charged metal plate (like charges repel and opposite charges attract)
  - Thomson concluded that the beam must be made of negatively charged particles, smaller than atoms.
  - since the atom was known to be electrically neutral, he developed a model in which negative charges were evenly scattered throughout an atom filled with a positively charged pudding like material
  - called the "Plum Pudding" model (or chocolate chip cookie dough)
- D. **Rutherford's Atomic Theory**
  - Rutherford discovered positively charged, fast moving particles called alpha particles
  - he shot some of these particles a piece of gold foil, thinking they would go straight through
  - when he actually performed the experiment, he found that the positively charged alpha particles were greatly deflected by the gold (figure 7, p. 104)
  - he determined that the alpha particles must have come very close to another positively charged particle causing them to be repelled and put off course
  - this led him to believe that the positive charge in an atom is not evenly distributed, but concentrated in a very small central area he called the nucleus
  - his model of the atom is a small, positively charged nucleus surrounded by negatively charged particles

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