

Acid-Base Titrations Prelab

1. Determine the volume released in the burets on the right. Remember significant figures and units!

Initial: 0.90 mL
Final: 20.21 mL
19.31 mL

2. If the concentration of the solution (NaOH) is 0.1500 M, calculate the number of moles of NaOH in the volume in (1). Show your work.

Moles of solute = concentration X volume
Volume used: final reading – initial reading
= 20.21 mL – 0.90 mL = 19.31 mL
Volume used = 19.31 mL / 1000 mL/L = 0.01931 L
Moles of NaOH: concentration X volume used
= 0.1500 M X 0.01931 L
= 0.0028965

Initial Reading



Final Reading



(For the following questions, please refer to the report on Blackboard.)

1. What is the concentration of the NaOH solution used in the titration?
Answer: 0.1500 M
2. What is the volume of the NaOH solution used in the titration?
Answer: 19.31 mL
3. What is the volume of the NaOH solution used in the titration?
Answer: 19.31 mL
4. What is the concentration of the NaOH solution used in the titration?
The initial concentration of the NaOH solution was 0.1500 M. The concentration of the NaOH solution used in the titration was 0.1500 M. The concentration of the NaOH solution used in the titration was 0.1500 M.

Chemfax Acid Base Titrations Lab Prelab Answers

Anthony S. Fauci



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