

Thermodynamic processes:

Process in which state of thermodynamic variables like P, V, T changes are called thermodynamic processes.

Given fixed number of moles / closed system ($n = \text{constant}$) under ideal conditions is assumed ($PV = nRT$ is applicable)

Constant volume process / Isochoric process / Isometric process:

Gas is confined in a piston cylinder arrangement having fixed piston.

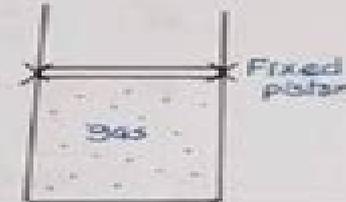
$$V = \text{constant}$$

$$PV = nRT \Rightarrow \frac{P}{T} = \left(\frac{nR}{V}\right) = \text{const.}$$

Process equation: $\frac{P}{T} = \text{constant}$

Between two states (P_1, V, T_1) and (P_2, V, T_2)

one can use: $\frac{P_1}{T_1} = \frac{P_2}{T_2}$ { T_1, T_2 must be in kelvin }



Example: A given mass of an ideal gas is confined in a vessel of fixed volume. At 27°C pressure of the gas is found to be 2.4 atm . Find its pressure if the gas is heated to 52°C .

Sol: $P_1 = 2.4 \text{ atm}$, $T_1 = 273 + 27 = 300 \text{ K}$
 $P_2 = ?$, $T_2 = 273 + 52 = 325 \text{ K}$
 $\frac{P_1}{T_1} = \frac{P_2}{T_2} \Rightarrow P_2 = \frac{T_2}{T_1} \times P_1 = \frac{325}{300} \times 2.4 = 2.6 \text{ atm}$ (answer)

Chemical Engineering Thermodynamics Lecture Notes

L Darling-Hammond



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