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Chemical Kinetics

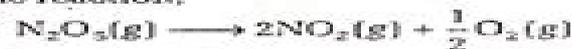
1. What is the activation energy for a reaction if its rate doubles when the temperature is raised from 20°C to 35°C?

($R = 8.314 \text{ J mol}^{-1} \text{ K}^{-1}$)

[2013]

- (a) 342 kJ mol⁻¹ (b) 269 kJ mol⁻¹
(c) 34.7 kJ mol⁻¹ (d) 15.1 kJ mol⁻¹
2. In a reaction, $A + B \rightarrow \text{Product}$, rate is doubled when the concentration of B is doubled and rate increases by a factor of 8 when the concentrations of both the reactants (A and B) are doubled. Rate law for the reaction can be written as [2012]
- (a) rate = $k[A][B]^2$
(b) rate = $k[A]^2[B]^2$
(c) rate = $k[A][B]$
(d) rate = $k[A]^2[B]$
3. In a zero order reaction for every 10°C rise of temperature, the rate is doubled. If the temperature is increased from 10°C to 100°C, the rate of the reaction will become [2013]
- (a) 256 times (b) 512 times
(c) 64 times (d) 128 times
4. Which one of the following statements for the order of a reaction is incorrect? [2011]
- (a) Order is not influenced by stoichiometric coefficient of the reactants
(b) Order of reaction is sum of power to the concentration terms of reactants to express the rate of reaction
(c) Order of reaction is always whole number
(d) Order can be determined only experimentally

5. For the reaction,



The value of rate of disappearance of N_2O_5 is given as $6.25 \times 10^{-3} \text{ mol L}^{-1} \text{ s}^{-1}$. The rate of formation of NO_2 and O_2 is given respectively as [2010]

- (a) $6.25 \times 10^{-3} \text{ mol L}^{-1} \text{ s}^{-1}$ and $6.25 \times 10^{-3} \text{ mol L}^{-1} \text{ s}^{-1}$
(b) $1.25 \times 10^{-2} \text{ mol L}^{-1} \text{ s}^{-1}$ and $3.125 \times 10^{-3} \text{ mol L}^{-1} \text{ s}^{-1}$
(c) $6.25 \times 10^{-3} \text{ mol L}^{-1} \text{ s}^{-1}$ and $3.125 \times 10^{-3} \text{ mol L}^{-1} \text{ s}^{-1}$
(d) $1.25 \times 10^{-2} \text{ mol L}^{-1} \text{ s}^{-1}$ and $6.25 \times 10^{-3} \text{ mol L}^{-1} \text{ s}^{-1}$
6. For an endothermic reaction, energy of activation is E_a and enthalpy of reaction is ΔH (both of these in kJ/mol). Minimum value of E_a will be [2010]
- (a) less than ΔH (b) equal to ΔH
(c) more than ΔH (d) equal to zero

7. During the kinetic study of the reaction, $2A + B \rightarrow C + D$, following results were obtained

Run	A/mol L ⁻¹	B/mol L ⁻¹	Initial rate of formation of D/mol L ⁻¹ min ⁻¹
I	0.1	0.1	6.0×10^{-3}
II	0.3	0.2	7.2×10^{-2}
III	0.3	0.4	2.88×10^{-1}
IV	0.4	0.1	2.40×10^{-2}

Based on the above data which one of the following is correct? [2010]

- (a) Rate = $k[A]^2[B]$ (b) Rate = $k[A][B]$
(c) Rate = $k[A]^2[B]^2$ (d) Rate = $k[A][B]^2$

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