

## Assessment 3

**Chapter Test B****Chapter: Chemical Equations and Reactions**

**Part 1** In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

- \_\_\_\_\_ 1. The production of a slightly soluble solid compound by a double-displacement reaction is called the formation of a
- gas.
  - precipitate.
  - redox reaction.
  - hydrogen.
- \_\_\_\_\_ 2. To balance a chemical equation, it is permissible to adjust the
- coefficients.
  - subscripts.
  - formula of the products.
  - number of products.
- \_\_\_\_\_ 3. In a chemical equation, the symbol (aq) indicates that the substance is
- solid.
  - dissolved in water.
  - an acid.
  - aqueous.
- \_\_\_\_\_ 4. The tendency for a replacement reaction to occur depends on the
- activity between any two elements in the activity series themselves.
  - temperature difference.
  - valence electrons and mass of.
  - activity between any two elements in the activity series themselves.
- \_\_\_\_\_ 5. The coefficients in a chemical equation
- indicate the number of moles of each substance that react.
  - show the number of grams of each substance that react.
  - are the molar masses of the substances.
  - show the valence electrons for each atom.
- \_\_\_\_\_ 6. If metal X is more reactive than metal Y in the activity series, then
- X will replace ions of Y in solution.
  - Y will replace ions of X in solution.
  - Y will form oxides only indirectly.
  - X will react with cold water.

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**Victor M. Corman**



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