

## MULTIPLE CHOICE QUESTIONS

1. When potassium dichromate crystals are heated with hydrochloric acid, we get  
(A) Oxygen (B) Hydrogen (C) Sulphur dioxide  
(D) Hydrogen sulphide (E) Chlorine Ans : E
2. Which of the following elements has lowest melting point?  
(A) Copper (B) Niobium (C) Molybdenum  
(D) Tungsten (E) Mercury Ans : E
3. In acidic medium if potassium permanganate is used as an oxidising agent, we get  
(A)  $MnO_4^{2-}$  (B)  $Mn^{2+}$  (C)  $MnO_2$   
(D)  $Mn_2O_3$  (E) None of these Ans : B
4. Among the following outermost configurations of transition metals, which shows the highest oxidation state?  
(A)  $3d^3 4s^2$  (B)  $3d^5 4s^1$  (C)  $3d^4 4s^2$   
(D)  $3d^6 4s^2$  (E)  $3d^7 4s^2$  Ans : C
5. Highest (+7) oxidation state is shown by  
(A) Co (B) Cr (C) V  
(D) Mn (E) Ni Ans : D
6. Potassium permanganate on heating gives  
(A)  $K_2MnO_4, MnO_2, O_2$  (B)  $MnO_2, O_2, KOH$  (C)  $K_2MnO_4, O_2$

# Chemistry Multiple Choice Questions With Answers

**LP Steffe**



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