

Chemistry B11
Chapters 10-13
Alkanes, Alkenes, Alkynes and Benzene

Organic compounds: organic chemistry is the chemistry of carbon and only a few other elements-chiefly, hydrogen, oxygen, nitrogen, sulfur, halogens, and phosphorus (from 116 elements).

Note: In the early days of chemistry, scientists thought organic compounds were those produced by living organisms and they could not synthesize any organic compound by starting with only inorganic compounds. In 1828, Friedrich Wöhler synthesized the first organic compound in his laboratory (urea).



Today, chemists obtain organic compounds in two principal ways: 1. Isolation from nature and 2. Synthesis in the laboratory.

Note: Compounds made in the laboratory are identical in both chemical and physical properties to those found in nature.

Why organic chemistry: we can find organic compounds everywhere around us (foods, flavors, fragrances, medicines, toiletries, cosmetics, plastic, paints, our body, and etc.). Chemist have discovered or synthesized more than 10 million of organic compounds. However, 1.7 million inorganic compounds are discovered or synthesized (85% of all known compounds are organic compounds).

Typical properties of organic compounds: 1. They contain carbon atom. 2. Bonding is almost entirely covalent (covalent compounds). 3. They have low boiling points and low melting points. 4. They are flammable (almost all burn). 5. They are soluble in nonpolar compounds and most are insoluble in water. 6. Many are gases, liquids, or solids.

It is important to know:

- Carbon:** normally forms four covalent bonds and has no unshared pairs of electrons.
- Hydrogen:** forms one covalent bond and has no unshared pairs of electrons.
- Nitrogen:** normally forms three covalent bonds and has one unshared pair of electrons.
- Oxygen:** normally forms two covalent bonds and has two unshared pairs of electrons.
- A Halogen:** normally forms one covalent bond and has three unshared pairs of electrons.

Functional group: an atom or group of atoms within a molecule that shows a characteristic set of predictable physical and chemical behaviors.

Functional groups are important in organic chemistry because:

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molecules illustrating the structural bonding involved and demonstrating the influence of these factors on fuel properties and reactions Engineering Thermodynamics details the construction of cyclic processes that can be followed by gases in order to accept heat and then convert a substantial proportion of it to mechanical work Combustion Principles examines the principles underpinning combustion processes and the concept of flammability Conventional Fuels discusses conventional fuels such as gasoline for spark ignition piston engines kerosine for gas turbine engines in large aircraft gas oil for high speed compression ignition piston engines diesel fuel and residual fuel for low speed compression ignition engines in ships and for the generation of electrical power and coal in regard to its use in power stations and its potential for conversion to liquids Combustion Performance draws together the main features of the first three basic chapters by illustrating the combustion performance of fuels burnt in heat engines Alternative Fuels examines the remaining lifetimes of the major hydrocarbon sources such as petroleum natural gas and coal It also outlines the manufacture properties and performance of various alternative fuels Overview of Conventional Fuels provides a review of the conventional fuels used in various sectors of light vehicle transport motor racing aerospace fleet operation and rail transport and the marine world Global Issues Affecting Transport presents methods of energy prediction that reflect the dynamics of global supply and the influence of new technologies Transport Fuels Technology Mobility for the Millennium **INIS Atomindex** ,1977

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