

**Mole Calculations Review Worksheet – answers on next page.**

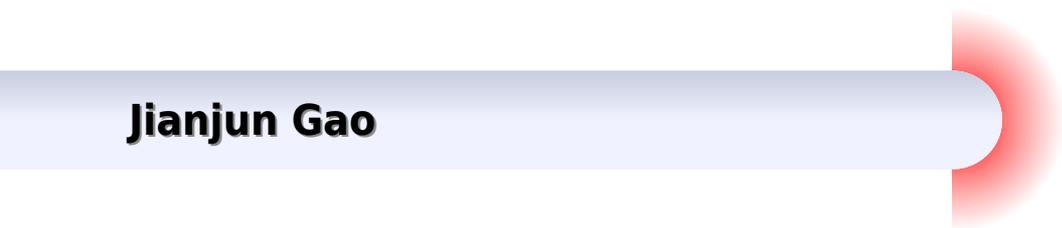
Complete the problems on this worksheet to review for the Chapter 10 test. Be sure to use dimensional analysis, and label all numbers with units when solving the problems.

- Calculate the molar mass of each compound.
  - LiOH
  - barium bromide
  - $\text{Mg}(\text{C}_2\text{H}_3\text{O}_2)_2$
  - $\text{Ca}(\text{NO}_3)_2$
- How many molecules are in 45.0 grams  $\text{CH}_4$ ?
- How many moles are in 18.8 grams NaOH?
- A salt shaker containing  $9.58 \times 10^{23}$  formula units NaCl contains how many moles?
- Which of the following has the greatest mass?  
4.2 moles C    or     $8.34 \times 10^{28}$  f.u.  $\text{CaCO}_3$     or    12.6 g  $\text{Al}(\text{NO}_3)_3$
- Find the number of representative particles in each of the following:
  - 0.28 mol Ni
  - 1.84 mol S
  - 1.32 mol  $\text{F}_2$
  - 3.15 mol  $\text{K}_3\text{PO}_4$
- How many liters would 1.45 moles oxygen gas occupy at STP?
- How many grams would 15.8 L of ammonia occupy at STP?
- How many total atoms are in 5.8 mol  $\text{CaCl}_2$ ?
- Calculate the percent composition of each element in the following compounds:
  - iron(III) oxide
  - $\text{Ca}_3(\text{PO}_4)_2$
  - methane
  - $\text{H}_2\text{SO}_4$
- Calculate the empirical and molecular formulas of adrenaline. The compound has a molecular mass of 183 g/mol, and is 59.0% C, 7.1% H, 26.2% O, and 7.7% N by mass.
- A compound is 72.2% Mg and 27.8% N by mass. Calculate the empirical formula.
- Nicotine is 74.1% C, 8.6% H, and 17.3% N by mass. The molecular mass of nicotine is 162 g/mol. Calculate its empirical and molecular formulas.
- How many total ions are in 2.95 grams  $\text{MgSO}_4$ ?

# Chemistry Calculation Review Name Chem Worksheet 12

1

**Jianjun Gao**



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