

Chemistry Notes 14.3

Three Basic Assumptions of the kinetic theory:

- 1) Gas is composed of particles, usually molecules or atoms
- 2) Particles in a gas move rapidly in constant random motion (they move in straight line paths only colliding with one another or other objects)
- 3) Collisions are perfectly elastic: kinetic energy is transferred without loss from one particle to another

Equation: $P \times V = n \times R \times T$

$R = 8.31 \text{ (L} \times \text{kPa) / (mol} \times \text{K)}$

$n = (P \times V) / (R \times T)$

Ideal Gas:

- 1) A gas is composed of particles with a mass that is low enough to be considered almost zero.
 - But a real gas has limitations to the pressure it can be compressed to before it condenses into a liquid

Real gases differ from ideal gases in the following ways:

- Real gases have small attractive and repulsive forces between particles, and ideal gases do not
- Real gas particles have a volume, and ideal gas particles do not
- Real gas particles collide inelastically and lose energy with collisions, and ideal gas particles collide elastically
- Real gases are not infinitely compressible

Ideal gases don't exist because they do take up space, and there are attractive forces

Diffusion: molecules moving from areas of high concentration to low concentration

- Perfume

Effusion: Gas escaping through tiny hole in container

- Both of these depend on the molar mass of the particle, which determines the speed
- The rate of effusion and diffusion is inversely proportional to the square root of the molar mass

To solve Graham's law problems, put whatever you're looking for in the numerator
6777

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