

## 10

## CHEMICAL QUANTITIES

## Practice Problems

On a piece of paper, solve the following problems.

**YOU MUST SHOW YOUR WORK TO GET CREDIT FOR YOUR ANSWERS!!!**

## SECTION 10.1 THE MOLE: A MEASUREMENT OF MATTER

1. What is the molar mass of sucrose ( $C_{12}H_{22}O_{11}$ )?
2. What is the molar mass of each of the following compounds?
  - a. phosphorus pentachloride ( $PCl_5$ )
  - b. uranium hexafluoride ( $UF_6$ )
3. Calculate the molar mass of each of the following ionic compounds:
  - a.  $KMnO_4$
  - b.  $Ca_3(PO_4)_2$
4. How many moles is  $3.52 \times 10^{24}$  molecules of water?
5. How many atoms of zinc are in 0.60 mol of zinc?
6. What is the mass of 1.00 mol of oxygen ( $O_2$ )?

## SECTION 10.2 MOLE–MASS AND MOLE–VOLUME RELATIONSHIPS

1. What is the molar mass of each of the following compounds?
  - a.  $C_6H_{12}O_6$
  - b.  $NaHCO_3$
  - c.  $C_2H_{12}$
  - d.  $KNH_4SO_4$
2. Calculate the mass in grams of each of the following:
  - a. 8.0 mol lead oxide ( $PbO$ )
  - b. 0.75 mol hydrogen sulfide ( $H_2S$ )
  - c. 0.00100 mol silicon tetrahydride ( $SiH_4$ )
  - d.  $1.50 \times 10^{-2}$  mol molecular oxygen ( $O_2$ )
  - e. 2.30 mol ethylene glycol ( $C_2H_6O_2$ )
3. How many grams are in 1.73 mol of dinitrogen pentoxide ( $N_2O_5$ )?
4. How many grams are in 0.658 mol of calcium phosphate [ $Ca_3(PO_4)_2$ ]?
5. Calculate the number of moles in each of the following:
  - a. 0.50 g sodium bromide ( $NaBr$ )
  - b. 13.5 g magnesium nitrate [ $Mg(NO_3)_2$ ]
  - c. 1.02 g magnesium chloride ( $MgCl_2$ )
  - d. 0.00100 g monochloromethane ( $CH_3Cl$ )
  - e.  $1.50 \times 10^{-3}$  g propylene glycol [ $C_3H_8(OH)_2$ ]
6. A chemist plans to use 435.0 grams of ammonium nitrate ( $NH_4NO_3$ ) in a reaction. How many moles of the compound is this?
7. A solution is to be prepared in a laboratory. The solution requires 0.0465 mol of quinine ( $C_{20}H_{24}N_2O_2$ ). What mass, in grams, should the laboratory technician obtain in order to make the solution?
8. What is the volume at STP of 2.66 mol of methane ( $CH_4$ ) gas?
9. How many moles is 135 L of ammonia ( $NH_3$ ) gas at STP?

# Chemistry Chapter 10 The Mole Study Answers

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