

AP Chemistry Final Study Guide

Unit 1: Atomic Structure and Properties (7%–9%)

- At STP, one mole of gas always occupies 22.4 L
- **Mass Spectroscopy:**
 - Atoms have different isotopes (carbon-12, carbon-14, etc) and this will affect the average atomic mass
 - The most abundant isotope will have the greatest effect
- **Coulomb's Law:**
 - The greater the charge of the nucleus, the more energy is required to ionize an electron
 - The closer an electron is to the nucleus, the more energy is required to ionize an electron
- **Electromagnetic Radiation:**
 - When atoms absorb energy, electrons jump to higher energy levels
 - When electrons drop from higher to lower energy levels, atoms give off electromagnetic radiation energy
- **Kinetic Energy of Ionized Electrons:**
 - The faster an ejected electron is going, the more kinetic energy it has
 - Electrons originally further away from the nucleus require less energy to eject, so it will have more kinetic energy
- **Electron Configuration:**
 - When entering the d-block (transition metals), the energy level drops by 1 unless you are using shorthand notation
 - Nickel: $1s^2 2s^2 2p^6 3s^2 3p^6 3d^8 4s^2$ (notice the d-block)
 - Nickel (with shorthand): $[\text{Ar}]4s^2 3d^8$
- **Law of Definite Proportions:**
 - A pure chemical compound broken down into elements always contains elements of a fixed ratio, independent of where and how it was created
- **Law of Multiple Proportions:**
 - When elements form compounds, the proportions of the elements in those chemical compounds can be expressed in small whole number ratios
- **Aufbau Principle:**
 - When building up the electron configuration of an atom, electrons are placed in orbitals, subshells, and shells in order of increasing energy

Chemistry Final Exam Study Guide

Scott C. Dulebohn



Chemistry Final Exam Study Guide:

Crimes Against the English Language , **CliffsNotes AP Chemistry** Bobrow Test Preparation Services,2009-02-09
The book itself contains chapter length subject reviews on every subject tested on the AP Chemistry exam as well as both sample multiple choice and free response questions at each chapter s end Two full length practice tests with detailed answer explanations are included in the book *Complete Book of Graduate Programs in the Arts and Sciences* Princeton Review (Firm),2004-09 Our Best 357 Colleges is the best selling college guide on the market because it is the voice of the students Now we let graduate students speak for themselves too in these brand new guides for selecting the ideal business law medical or arts and humanities graduate school It includes detailed profiles rankings based on student surveys like those made popular by our Best 357 Colleges guide as well as student quotes about classes professors the social scene and more Plus we cover the ins and outs of admissions and financial aid Each guide also includes an index of all schools with the most pertinent facts such as contact information And we ve topped it all off with our school says section where participating schools can talk back by providing their own profiles It s a whole new way to find the perfect match in a graduate school

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The Publishers' Trade List Annual, 1978

Chemistry Steven S. Zumdahl, 1993

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