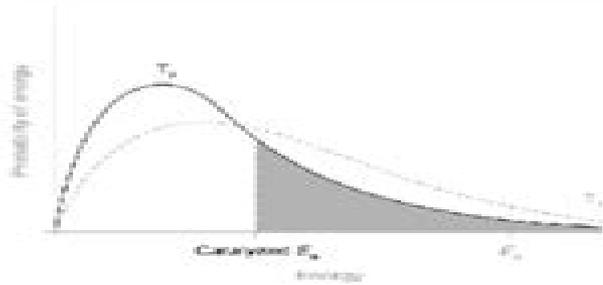


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2.	f	<p>a. Zn^{2+} ions in full shell OR dentic number is zero since no unpaired d orbitals</p>	<p>Do not accept Zn in any transition metal. Do not accept alternative for Zn^{2+} ions.</p>	1
3.	b	<p>ligands donate pairs of electrons to metal ions OR form coordinate covalent bonds</p> <p>ligands use lone pairs AND metal uses empty d orbitals</p>		1
4.	d	<p>Electrostatic attraction ✓ Between oppositely charged ions/between Fe^{2+} and SO_4^{2-} ✓</p>		2
4.	a	 <p>same higher AND to left of T_2 ✓ reactant E_a reduced AND to the left of E_a of curve T_1 ✓</p>	<p>Do not promote curve into a table, and putting labels (height and width, or colour) in case after it ✓</p> <p>Do not award AP if curve shown shows significantly more than reactant formation than reactant curve than curve 1.</p> <p>Accept E_a closer to T_2 instead of curve shown or closer to left of modified E_a</p>	2
5.	b	<p>to form AND react ✓</p>		1

3. Equilibrium favours backward direction/shifts to LHS **AND**

As increase in pressure favours side with less moles of gas.

[1]

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