

CHAPTER 24 REVIEW

Nuclear Chemistry

SECTION 1

Write a complete answer to the following questions in the space provided.

1. Describe the relationship between the three elementary particles in the table below from their properties.
 - a. Neutrons.
 - b. Electrons.
 - c. Positrons.
 - d. They all have the same mass.
2. The electron-positron annihilation reaction is:
 - a. an energy-releasing event.
 - b. an energy-absorbing event.
 - c. an exothermic reaction.
 - d. an endothermic event.
3. The number of electrons in each atom of the:
 - a. number of protons.
 - b. number of neutrons.
 - c. number of electrons.
 - d. ratio of neutrons to protons.
4. If an atom should lose three protons, the atom's:
 - a. number of electrons will increase.
 - b. mass is increased threefold.
 - c. number of protons will increase.
 - d. number of protons will decrease.
5. The nature of a gamma-ray emission from a radioactive substance is:
 - a. similar to alpha rays, but less massive.
 - b. gamma rays, which are less massive.
 - c. different from alpha rays, but have the same mass as alpha rays.
 - d. similar to alpha rays, but more massive.
6. Based on Figure 24.1 of the text, which range of β^- has the greatest:
 - a. least the smaller binding energy per nucleon?
 - b. greatest number of nucleons per nucleon?
7. The number of nucleons in an atom of hydrogen-2 is _____.
8. The mass of that atom is _____.

Chapter 21 Nuclear Chemistry Section 3 Answers

Kathleen Armour



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Radiochemistry and Nuclear Chemistry Gregory Choppin, Jan-Olov Liljenzin, Jan Rydberg, Christian Ekberg, 2013-09-05 Radiochemistry or nuclear chemistry is the study of radiation from an atomic and molecular perspective including elemental transformation and reaction effects as well as physical health and medical properties This revised edition of one of the earliest and best known books on the subject has been updated to bring into teaching the latest developments in research and the current hot topics in the field To further enhance the functionality of this text the authors have added numerous teaching aids examples in MathCAD with variable quantities and options hotlinks to relevant text sections from the book and online self grading tests New edition of a well known respected text in the specialized field of nuclear radiochemistry Includes an interactive website with testing and evaluation modules based on exercises in the book Suitable for both radiochemistry and nuclear chemistry courses

Chemistry John A. Olmsted, Gregory M. Williams, Robert Charles Burk, 2016-01-14 Olmsted Burk is an introductory general chemistry text designed specifically with Canadian professors and students in mind A reorganized Table of Contents and inclusion of SI units IUPAC standards and Canadian content designed to engage and motivate readers distinguish this text from many of the current text offerings It more accurately reflects the curriculum of most Canadian institutions Instructors will find the text sufficiently rigorous while it engages and retains student interest through its accessible language and clear problem solving program without an excess of material that makes most text appear daunting and redundant

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Concepts in Action Program Planner National Chemistry Physics Earth Science ,2003-11 Prentice Hall Physical Science Concepts in Action helps students make the important connection between the science they read and what they experience every day Relevant content lively explorations and a wealth of hands on activities take students understanding of science beyond the page and into the world around them Now includes even more technology tools and activities to support differentiated instruction *Understanding Physical Chemistry* Arthur W. Adamson,1980 **Study Guide with Solutions for Organic Chemistry, Third Edition** Ralph J. Fessenden,Joan S. Fessenden,1986 Provides a set of additional drill problems chapter by chapter discussions and supplemental instructional material to help students master organic chemistry problem solving techniques *Student Solutions Guide to Accompany Introduction to Organic Chemistry, Second Edition*, by William H. Brown David R. Benson,William H. Brown,Brent L. Iverson,Sheila A. Iverson,2000 *Chemistry* John S. Phillips,Cheryl Wistrom,2000 *Chemistry for Health-related Sciences* Curtis T. Sears,Conrad L. Stanitski,1983 **AP Chemistry** Albert S. Tarendash,Frederick J. Rowe,2002 This indispensable guide to chemistry helps students who wish to prepare for the AP Chemistry exam on their own Comprehensive and easy to understand this learning guide includes a full content review two full length practice tests with hundreds of practice questions and thorough answer explanations and proven test taking strategies Chemistry James C. Hill,2000 This text defines the concepts needed to learn or review cardiac auscultation The combination of audio and text explains how to identify and interpret normal and common abnormal heart sounds Some heart sounds are reproduced on a heart sound simulator allowing for a clear crisp grasp of specific individual sounds Others are recorded from real patients to distinguish between similar heart and lung sounds and to help the listener select the heart sounds from the auditory milieu **Study Guide for Chemistry, Third Edition [by] Steven S. Zumdahl** Paul B. Kelter,1993 **The British National Bibliography Cumulated Subject Catalogue** ,1960 CHEMISTRY SILBERBERG,2003 Physical Chemistry Thomas Engel,2006 *Chemical Principles* Steven S. Zumdahl,1992 The Study Guide reflects the unique problem solving approach taken by the Chemical Principles text The new edition of the Study Guide includes many new worked out examples

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