

Directions: Read carefully and solve the following problems.

1. The complete combustion of Octane results to the production of Carbon dioxide and water as shown in the unbalanced chemical equation below.



- a. Write the balanced chemical reaction equation.
 - b. How much CO_2 in grams will be produced from burning 30 grams of Octane?
 - c. What mass of Oxygen is needed to completely consume 100 grams of Octane?
2. In a certain experiment, Hydrochloric acid is used to decompose Calcium carbonate. The chemical reaction equation is shown below.



- a. After the experiment, the total mass of $CaCl_2$ produced is 120g. Assuming that the reaction proceeded to completion, how many moles of $CaCO_3$ were used in the experiment?
 - b. Based on the amount of $CaCl_2$ produced in the reaction, how much CO_2 in grams effervesced out of the Calcium carbonate used?
3. Hydrogen fluoride is produced by reacting Calcium fluoride with sulfuric acid. A by-product of the process is Calcium sulfate.

Chapter 3 Stoichiometry Exercises Answers

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