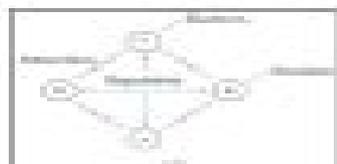


Chapter 9 - Covalent Bonding

I. Covalent Bonding: attractive force produced as a result of shared electrons. (pp. 189-192)

A. A _____ is formed when two or more atoms bond covalently.

B. Bonds form when there is a balance of attractive and repulsive forces between two atoms.



C. Single Covalent Bond & Lewis Structures

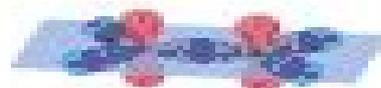
- Single Covalent Bond also called _____ bond [σ]
- Orbitals overlap end to end, concentrating the electrons in a _____ between the two atoms
- One pair of electrons is shared between the two atoms

D. Multiple Covalent Bond & Lewis Structures

- Sharing _____ than _____ of electrons
- Carbon, nitrogen, oxygen, and sulfur most often form _____ bonds
- Shared pairs may be needed to form a complete _____
- Double bond = two pairs of electrons shared For example: $O=O$
- Triple bond = three pairs of electrons shared For example: $N\equiv N$

The pi bond (π) forms when parallel p orbitals overlap.

A pi bond always accompanies a sigma bond.



This diagram shows how DOUBLE bonds form. First, a sigma (σ) bond forms (electron clouds overlap end-to-end). Then, two p orbitals overlap side-to-side to form a pi (π) bond.

These diagrams show the same molecule above simplified. That strange looking shape is what a pi bond actually looks like!



Chapter 6 Covalent Bonding

G Psacharopoulos



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