

CHAPTER 6

Thermochemistry

CHAPTER TERMS AND DEFINITIONS

Numbers in parentheses after definitions give the text sections in which the terms are explained. Starred terms are italicized in the text. Where a term does not fall directly under a text section heading, additional information is given for you to locate it.

thermodynamics* science of the relationships between heat and other forms of energy (6.1, introductory section)

thermochemistry* one area of thermodynamics; study of the quantity of heat absorbed or evolved by chemical reactions (6.1, introductory section)

energy potential or capacity to move matter (6.1)

kinetic energy (E_k) energy associated with an object by virtue of its motion (6.1)

joule (J) SI unit of energy, $\text{kg} \cdot \text{m}^2/\text{s}^2$ (6.1)

watt* measure of quantity of energy used per unit time; 1 J/s (6.1)

calorie (cal) non-SI unit of energy commonly used by chemists; originally defined as the amount of energy required to raise the temperature of one gram of water by one degree Celsius; $1 \text{ cal} = 4.184 \text{ J}$ (exact definition) (6.1)

potential energy (E_p) an object's energy because of its position in a field of force (6.1)

internal energy (U) sum of kinetic and potential energies of the particles making up a substance (6.1)

E_{tot} * total energy of a substance; sum of the kinetic, potential, and internal energies of the substance (6.1)

law of conservation of energy energy may be converted from one form to another, but the total quantity of energy remains constant (6.1)

thermodynamic system (or system) substance or mixture of substances under study in which a physical or chemical change occurs (6.2)

surroundings everything in the vicinity of a thermodynamic system (6.2)

heat (q) energy that flows between system and surroundings because of a difference in temperature between the thermodynamic system and its surroundings (6.2)

thermal equilibrium* state in which energy does not flow as heat between system and surroundings; temperature equality (6.2)

heat of reaction value of q required, at a given temperature, to return a system to the given temperature at the completion of the reaction (6.2)

exothermic process chemical reaction or physical change in which heat is evolved (q is negative) (6.2)

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Minjie Lin



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