

CHAPTER 6 REVIEW*Chemical Bonding***MIXED REVIEW**

SHORT ANSWER Answer the following questions in the space provided.

1. Name the type of energy that is a measure of strength for each of the following types of bonds:

_____ **lattice energy** _____ a. ionic bond
 _____ **bond energy** _____ b. covalent bond
 _____ **enthalpy of vaporization** _____ c. metallic bond

2. Use the electronegativity values shown in **Figure 20**, on page 161 of the text, to determine whether each of the following bonds is nonpolar covalent, polar covalent, or ionic.

_____ **ionic** _____ a. H—F _____ **nonpolar covalent** _____ d. H—H
 _____ **ionic** _____ b. Na—Cl _____ **polar covalent** _____ e. H—C
 _____ **polar covalent** _____ c. H—O _____ **polar covalent** _____ f. H—N

3. How is a hydrogen bond different from an ionic or covalent bond?

A hydrogen bond is a dipole-dipole attraction between a partially positive hydrogen atom and the unshared electron pair of a strongly electronegative atom such as O, N, or F. Unlike ionic or covalent bonds, in which electrons are given up or shared, the hydrogen bond is a weaker attraction. Hydrogen bonds are generally intermolecular, while ionic and covalent bonds occur between ions or atoms respectively.

4. H₂S and H₂O have similar structures and their central atoms belong to the same group. Yet H₂S is a gas at room temperature and H₂O is a liquid. Use bonding principles to explain why this is.

Oxygen has higher electronegativity than sulfur, which creates a highly polar bond. Increased polarity in H₂O bonds means a stronger intermolecular attraction, making water a liquid at room temperature. Hydrogen bonding exists between water molecules, but not between hydrogen sulfide molecules.

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Liying Dong



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Biology Neil A. Campbell, Jane B. Reece, 2005 Neil Campbell and Jane Reece's BIOLOGY remains unsurpassed as the most successful majors biology textbook in the world. This text has invited more than 4 million students into the study of this dynamic and essential discipline. The authors have restructured each chapter around a conceptual framework of five or six big ideas. An Overview draws students in and sets the stage for the rest of the chapter; each numbered Concept Head announces the beginning of a new concept, and Concept Check questions at the end of each chapter encourage students to assess their mastery of a given concept. New Inquiry Figures focus students on the experimental process, and new Research Method Figures illustrate important techniques in biology. Each chapter ends with a Scientific Inquiry Question that asks students to apply scientific investigation skills to the content of the chapter.

Intermetallics Rainer Pöttgen, Dirk Johrendt, 2019-07-22 The expanded edition focuses still more on Synthesis, discussing necessary requirements for sample preparation and presents the broad range from structural analysis to property investigations. Additional examples of chemical and physical properties are highlighted for metallic binary and multinary intermetallic compounds. The work contains an updated literature overview in all sub-chapters and a detailed formulae index.

Journal of Chemical Research, 1977 Contains papers abstracted in Journal of Chemical Research Synopses.

International Tables for X-ray Crystallography: Physical and chemical tables, editors: C.H. MacGillavry and G.D. Rieck; general editor: K. Lonsdale International Union of Crystallography, 1962

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