

## Section 12.1

### Section 12.1

Answer complete answer multiple-choice questions that are provided.

1. \_\_\_\_\_
  - a. \_\_\_\_\_
    - (a) correct, in terms of all the reactants and products.
    - (b) correct, in terms of moles of reactants and products.
    - (c) correct, in terms of work done by the system on the surroundings.
    - (d) correct, in terms of volume change, but not in terms of work done.
  - b. \_\_\_\_\_
    - (a) the mass of the following reactants:  $\text{H}_2$ ,  $\text{O}_2$ , and  $\text{H}_2\text{O}$ .
    - (b) the mass of the following products:  $\text{H}_2\text{O}$  and  $\text{H}_2$ .
    - (c) the number of moles of  $\text{H}_2$  that will react with a certain quantity of  $\text{O}_2$ .
    - (d) the amount of energy released when the water is formed.
  - c. \_\_\_\_\_
    - (a) the number of moles of  $\text{H}_2$  that will react with a certain quantity of  $\text{O}_2$ .
    - (b) the energy released in the reaction.
    - (c) the molar enthalpies of all substances in the reaction.
    - (d) the number of moles of  $\text{H}_2$  that will react with a certain quantity of  $\text{O}_2$ .
  - d. \_\_\_\_\_
    - (a) the number of moles of  $\text{H}_2$  that will react with a certain quantity of  $\text{O}_2$ .
    - (b) the number of moles of  $\text{H}_2$  that will react with a certain quantity of  $\text{O}_2$ .
    - (c) the number of moles of  $\text{H}_2$  that will react with a certain quantity of  $\text{O}_2$ .
    - (d) the number of moles of  $\text{H}_2$  that will react with a certain quantity of  $\text{O}_2$ .

2. Answer the two questions posed by the following balanced equation:



a. Balance the equation:



b. What is the coefficient of  $\text{H}_2$  in the balanced equation?

c. How many moles of  $\text{H}_2$  are required to produce 1.0 mole of  $\text{H}_2\text{O}$ ?

d. How many moles of  $\text{O}_2$  are required to produce 1.0 mole of  $\text{H}_2\text{O}$ ?

e. How many moles of  $\text{H}_2$  are required to produce 1.0 mole of  $\text{H}_2\text{O}$ ?

# Chapter 9 Review Stoichiometry Section 1 Answers

**SJ Ball**



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