

## Section 1: Chemistry

### Section 1: Chemistry

Section 1: Chemistry

1. A. The following is a chemical equation representing the reaction of hydrogen gas and oxygen gas to form water.
  - (a) Identify the products of this reaction.
  - (b) Identify the reactants of this reaction.
  - (c) Balance the equation.
2. B. The following is a chemical equation representing the reaction of hydrogen gas and oxygen gas to form water.
  - (a) Identify the products of this reaction.
  - (b) Identify the reactants of this reaction.
  - (c) Balance the equation.
3. C. The following is a chemical equation representing the reaction of hydrogen gas and oxygen gas to form water.
  - (a) Identify the products of this reaction.
  - (b) Identify the reactants of this reaction.
  - (c) Balance the equation.
4. D. The following is a chemical equation representing the reaction of hydrogen gas and oxygen gas to form water.
  - (a) Identify the products of this reaction.
  - (b) Identify the reactants of this reaction.
  - (c) Balance the equation.

Section 2: Chemistry

$$2H_2 + O_2 \rightarrow 2H_2O$$

a. Balance the equation.

$$2H_2 + O_2 \rightarrow 2H_2O$$

b. How many moles of  $H_2$  are needed to react with 1 mole of  $O_2$ ?

c. How many moles of  $H_2O$  are produced when 2 moles of  $H_2$  react with 1 mole of  $O_2$ ?

d. How many moles of  $H_2$  are needed to react with 1 mole of  $O_2$  to produce 2 moles of  $H_2O$ ?

e. How many moles of  $H_2O$  are produced when 2 moles of  $H_2$  react with 1 mole of  $O_2$ ?

# Chapter 9 Review Stoichiometry Section 3

**Arturo Cuomo**



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