

## Section 18.4 Neutralization

In your textbook, read about neutralization and titration.

For each item in Column A, write the letter of the matching item in Column B.

### Column A

- \_\_\_\_\_ 1. A chemical dye that changes color based on the pH of a solution
- \_\_\_\_\_ 2. A method for using a neutralization reaction to determine the concentration of a solution
- \_\_\_\_\_ 3. A reaction in which an acid and a base react to produce a salt and water
- \_\_\_\_\_ 4. A solution of known concentration
- \_\_\_\_\_ 5. An ionic product of an acid-base reaction
- \_\_\_\_\_ 6. The point in a titration in which an indicator changes color
- \_\_\_\_\_ 7. The stoichiometric point of a titration

### Column B

- a. acid-base indicator
- b. end point
- c. equivalence point
- d. neutralization
- e. salt
- f. standard solution
- g. titration

Complete the following table, indicating the formula and name of the salt formed by a neutralization reaction between the listed acid and base.

Acid	Base	Salt formula	Salt name
8. HCl	KOH	KCl	potassium chloride
9. $H_2SO_4$	$Mg(OH)_2$		
10. $H_3PO_4$	NaOH		
11. $HNO_3$	$Fe(OH)_3$		
12. $H_3PO_4$	$Ca(OH)_2$		

In the space at the left, write 1 through 4 to show the correct sequence of the steps in performing a titration using a pH meter. Then, write 5 through 8 to sequence the steps used to calculate the concentration of the unknown solution.

### Sequence of Steps

- \_\_\_\_\_ 13. Continue adding the standard solution to the solution of unknown concentration until the equivalence point is reached.
- \_\_\_\_\_ 14. Fill a buret with the standard solution.
- \_\_\_\_\_ 15. Start adding the standard solution slowly, with mixing, to the solution of unknown concentration, reading the pH at regular intervals.

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## **Table of Contents Chapter 9 Section 1 Answer Key**

1. Understanding the eBook Chapter 9 Section 1 Answer Key
  - The Rise of Digital Reading Chapter 9 Section 1 Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 9 Section 1 Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chapter 9 Section 1 Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 9 Section 1 Answer Key
  - Personalized Recommendations

- Chapter 9 Section 1 Answer Key User Reviews and Ratings
- Chapter 9 Section 1 Answer Key and Bestseller Lists
- 5. Accessing Chapter 9 Section 1 Answer Key Free and Paid eBooks
  - Chapter 9 Section 1 Answer Key Public Domain eBooks
  - Chapter 9 Section 1 Answer Key eBook Subscription Services
  - Chapter 9 Section 1 Answer Key Budget-Friendly Options
- 6. Navigating Chapter 9 Section 1 Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Chapter 9 Section 1 Answer Key Compatibility with Devices
  - Chapter 9 Section 1 Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chapter 9 Section 1 Answer Key
  - Highlighting and Note-Taking Chapter 9 Section 1 Answer Key
  - Interactive Elements Chapter 9 Section 1 Answer Key
- 8. Staying Engaged with Chapter 9 Section 1 Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chapter 9 Section 1 Answer Key
- 9. Balancing eBooks and Physical Books Chapter 9 Section 1 Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chapter 9 Section 1 Answer Key
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Chapter 9 Section 1 Answer Key
  - Setting Reading Goals Chapter 9 Section 1 Answer Key
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chapter 9 Section 1 Answer Key
  - Fact-Checking eBook Content of Chapter 9 Section 1 Answer Key

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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