

WORKSHEET
CHEMISTRY FOR ENGINEERING STUDENTS, LECTURE 1
 Molecular formula of hydrocarbon or a mixture of hydrocarbons

ANSWERS KEY

QUESTION 1. Calculate the empirical formula of hydrocarbon and its molecular formula.

1. A hydrocarbon of which 0.1000 mole is completely burnt by 0.345 g of oxygen yields carbon dioxide and water respectively 0.1500 mol and 0.1500 mol . Find the empirical formula of hydrocarbon and its molecular formula.
 - (a) write the balanced chemical equation for the reaction
 $C_xH_y + O_2 \rightarrow CO_2 + H_2O$
 - (b) calculate the relative atomic weights of hydrocarbon as though the 0.1000 mol of hydrocarbon reacts with the given amount of oxygen and 0.345 g of oxygen. $(12x + 1.5y)$
 - (c) find the relative atomic weights of hydrocarbon as though the 0.1000 mol of hydrocarbon reacts with the given amount of oxygen and 0.345 g of oxygen. $(12x + 1.5y) = 3.45$

2. A hydrocarbon hydrocarbon contains 83.26% carbon and the relative atomic weight of hydrocarbon is 58. The hydrocarbon reacts with oxygen to form carbon dioxide and water vapor. These gases cooling of constant temperature are condensed by 1.00 g of water hydrocarbon.
 - (a) write the balanced chemical equation for the reaction
 $C_xH_y + O_2 \rightarrow CO_2 + H_2O$
 - (b) calculate the relative atomic weights of hydrocarbon as though 0.1000 mol of hydrocarbon reacts with the given amount of oxygen. $(12x + 1.5y)$
 - (c) calculate the relative atomic weights of hydrocarbon as though 0.1000 mol of hydrocarbon reacts with the given amount of oxygen. $(12x + 1.5y) = 58$

3. Hydrocarbon of which 0.1000 mole reacts with 0.345 g of oxygen to form carbon dioxide and water respectively 0.1500 mol and 0.1500 mol .
 - (a) write the balanced chemical equation for the reaction
 $C_xH_y + O_2 \rightarrow CO_2 + H_2O$
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 - (c) find the relative atomic weights of hydrocarbon as though 0.1000 mol of hydrocarbon reacts with the given amount of oxygen and 0.345 g of oxygen. $(12x + 1.5y) = 3.45$

QUESTION 2. Calculate the empirical formula of hydrocarbon.

1. A hydrocarbon of which 0.1000 mole is completely burnt by 0.345 g of oxygen yields carbon dioxide and water respectively 0.1500 mol and 0.1500 mol . Find the empirical formula of hydrocarbon and its molecular formula.
 - (a) write the balanced chemical equation for the reaction
 $C_xH_y + O_2 \rightarrow CO_2 + H_2O$
 - (b) calculate the relative atomic weights of hydrocarbon as though 0.1000 mol of hydrocarbon reacts with the given amount of oxygen and 0.345 g of oxygen. $(12x + 1.5y) = 3.45$

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